

## Chapter 4 and 5 Study Guide (Atomic Structure and the Periodic Table)

1. What did Dalton's model look like? List the parts of his atomic theory.

His model of an atom was a solid sphere.



Parts of the atomic theory:

- All matter is made up of individual particles called atoms, which can't be divided.
  - All elements made of atoms
  - Atoms of same element have same mass, atoms of different elements have different masses
  - Compounds have more than one type of element
  - Atoms of different elements always combine the same way in compound

2. What did JJ Thompson discover? Draw/describe his model.

JJ Thompson discovered the electron. "Plum pudding model" evenly scattered charges.

3. What did Millikan discover and draw/describe his model.

Millikan discovered the exact charge and mass of an electron. His model was of an atom was that there were negative charges scattered in a positive substance.

4. What did Rutherford discover? Draw/describe his model.

Discovered the dense, positive center of an atom. Model was that protons were in center, electrons moved on the outside.

5. What did Bohr discover? Draw/describe his model.

Electrons moved in energy levels. Nucleus at center, electrons moving in energy levels on the outside.

6. To find the atomic mass what two subatomic particles do you add together?

Protons + neutrons

7. Atomic number equals the number of \_\_\_\_\_ protons \_\_\_\_\_ and \_\_\_\_\_ electrons \_\_\_\_\_.

8. What are the three main parts to the atom? Name them, their charge, location and their mass.

Particle	Charge	Location	Mass
Electron	-	Outside nucleus	0 amu
Proton	+	Nucleus	1 amu
Neutron	0	Nucleus	1 amu

# Periodic Table Test Answer Key

**Todd Stanley, Betsy Moore**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author names.

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