

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|-----------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

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SA Adler



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Table of Contents Periodic Trends Lab Answer Key

1. Understanding the eBook Periodic Trends Lab Answer Key
 - The Rise of Digital Reading Periodic Trends Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Lab Answer Key
 - Personalized Recommendations
 - Periodic Trends Lab Answer Key User Reviews and Ratings
 - Periodic Trends Lab Answer Key and Bestseller Lists
5. Accessing Periodic Trends Lab Answer Key Free and Paid eBooks
 - Periodic Trends Lab Answer Key Public Domain eBooks
 - Periodic Trends Lab Answer Key eBook Subscription Services
 - Periodic Trends Lab Answer Key Budget-Friendly Options

6. Navigating Periodic Trends Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Lab Answer Key Compatibility with Devices
 - Periodic Trends Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Lab Answer Key
 - Highlighting and Note-Taking Periodic Trends Lab Answer Key
 - Interactive Elements Periodic Trends Lab Answer Key
8. Staying Engaged with Periodic Trends Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Lab Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Lab Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Lab Answer Key
 - Setting Reading Goals Periodic Trends Lab Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Lab Answer Key
 - Fact-Checking eBook Content of Periodic Trends Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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