

**6.3****PERIODIC TRENDS****Section Review****Objectives**

- Describe trends among elements for atomic size
- Explain how ions form
- Describe and explain periodic trends for first ionization energy, ionic size, and electronegativity

**Vocabulary**

- atomic radius
- ion
- cation
- anion
- ionization energy
- electronegativity

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Atomic radii generally \_\_\_\_\_1\_\_\_\_\_ as you move from left to right in a period. Atomic size \_\_\_\_\_2\_\_\_\_\_ with atomic number within a group because there are more occupied \_\_\_\_\_3\_\_\_\_\_ and an increased shielding effect, despite an increase in nuclear \_\_\_\_\_4\_\_\_\_\_.

The energy required to remove an electron from an atom is known as \_\_\_\_\_5\_\_\_\_\_ energy. This quantity generally \_\_\_\_\_6\_\_\_\_\_ as you move left to right across a period. Ions form when \_\_\_\_\_7\_\_\_\_\_ are transferred between atoms. Cations are always \_\_\_\_\_8\_\_\_\_\_ than the atoms from which they form. The ability of an atom to attract electrons when it is in a compound is called \_\_\_\_\_9\_\_\_\_\_, and this value \_\_\_\_\_10\_\_\_\_\_ as you move from left to right across a period.

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**Part B True-False**

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 11. Compounds are composed of particles called ions.



# Periodic Trends Section Review Answers

**Silberberg**





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**Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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### **Table of Contents Periodic Trends Section Review Answers**

1. Understanding the eBook Periodic Trends Section Review Answers
  - The Rise of Digital Reading Periodic Trends Section Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Section Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Section Review Answers
  - Personalized Recommendations
  - Periodic Trends Section Review Answers User Reviews and Ratings
  - Periodic Trends Section Review Answers and Bestseller Lists



5. Accessing Periodic Trends Section Review Answers Free and Paid eBooks
  - Periodic Trends Section Review Answers Public Domain eBooks
  - Periodic Trends Section Review Answers eBook Subscription Services
  - Periodic Trends Section Review Answers Budget-Friendly Options
6. Navigating Periodic Trends Section Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Section Review Answers Compatibility with Devices
  - Periodic Trends Section Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Section Review Answers
  - Highlighting and Note-Taking Periodic Trends Section Review Answers
  - Interactive Elements Periodic Trends Section Review Answers
8. Staying Engaged with Periodic Trends Section Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Section Review Answers
9. Balancing eBooks and Physical Books Periodic Trends Section Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Section Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Section Review Answers
  - Setting Reading Goals Periodic Trends Section Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Section Review Answers
  - Fact-Checking eBook Content of Periodic Trends Section Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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