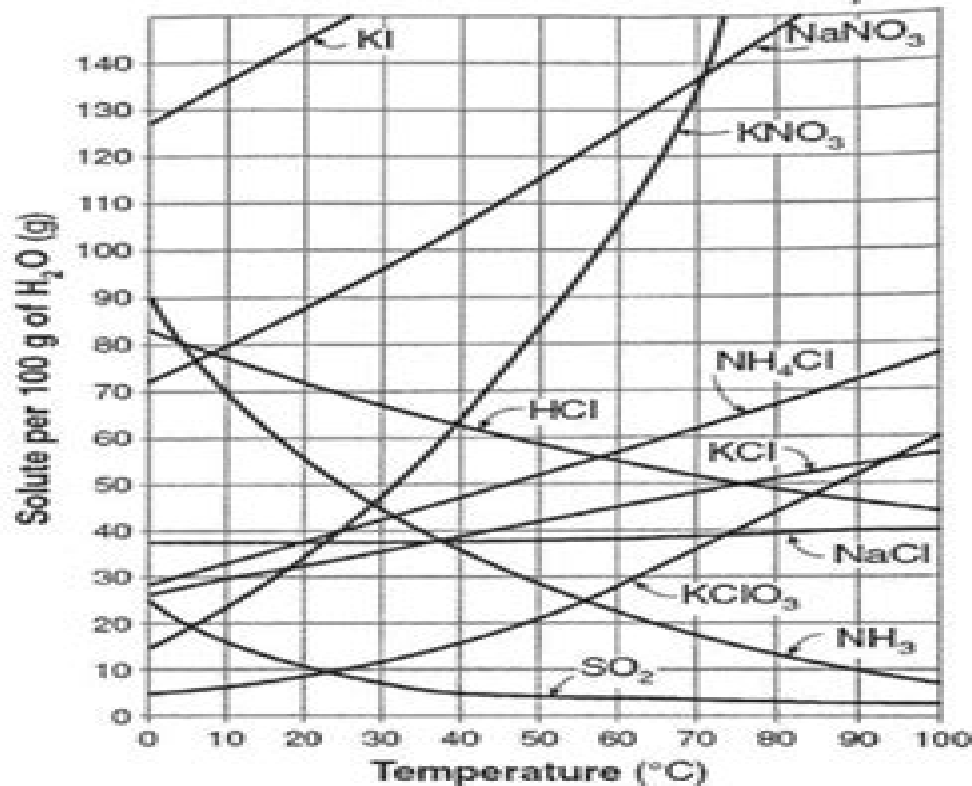


Reading a Solubility Chart

- 1) The curve shows the # of grams of solute in a saturated solution containing 100 mL or 100 g of water at a certain temperature.
- 2) Any amount of solute below the line indicates the solution is unsaturated at a certain temperature
- 3) Any amount of solute above the line in which all of the solute has dissolved shows the solution is supersaturated.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution - # g of a saturated solution at that temperature. (according to the curve)
- 5) Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



Solubility Problems

1. Which of the salts shown on the graph is the most soluble in water at 10°C? **KI**
2. At what temperature is 90g of NaNO₃ a saturated solution in 100grams of water? **≈ 23°C**
3. A solution has 70grams of KNO₃ and is unsaturated, what is the minimum temperature? **46°C**
4. At 20°C, a saturated solution of KClO₃ is created in water. How many grams of potassium chlorate must be added to saturate the solution at 50°C? **12g**
5. Which of the salts shown on the graph is the least soluble in water at 10°C? **KClO₃**
6. Which of the salts shown on the graph has the greatest increase in solubility as the temperature increases from 30 degrees to 60 degrees? **KNO₃**

Reading A Solubility Chart Answers

The Princeton Review



Reading A Solubility Chart Answers:

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Table of Contents Reading A Solubility Chart Answers

1. Understanding the eBook Reading A Solubility Chart Answers
 - The Rise of Digital Reading Reading A Solubility Chart Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reading A Solubility Chart Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reading A Solubility Chart Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reading A Solubility Chart Answers
 - Personalized Recommendations
 - Reading A Solubility Chart Answers User Reviews and Ratings
 - Reading A Solubility Chart Answers and Bestseller Lists

5. Accessing Reading A Solubility Chart Answers Free and Paid eBooks
 - Reading A Solubility Chart Answers Public Domain eBooks
 - Reading A Solubility Chart Answers eBook Subscription Services
 - Reading A Solubility Chart Answers Budget-Friendly Options
6. Navigating Reading A Solubility Chart Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reading A Solubility Chart Answers Compatibility with Devices
 - Reading A Solubility Chart Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reading A Solubility Chart Answers
 - Highlighting and Note-Taking Reading A Solubility Chart Answers
 - Interactive Elements Reading A Solubility Chart Answers
8. Staying Engaged with Reading A Solubility Chart Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reading A Solubility Chart Answers
9. Balancing eBooks and Physical Books Reading A Solubility Chart Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reading A Solubility Chart Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reading A Solubility Chart Answers
 - Setting Reading Goals Reading A Solubility Chart Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reading A Solubility Chart Answers
 - Fact-Checking eBook Content of Reading A Solubility Chart Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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