



Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = 30.0 \text{ g/mol}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

Water (H₂O) gmm:

$$(2)(1.0) + (1)(16.0) = 18.0 \text{ g/mol}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100 = 88.9\%$$

$$\frac{45 \text{ g NO} \times \frac{1 \text{ mol NO}}{30 \text{ g NO}} \times \frac{4 \text{ mol NO}}{4 \text{ mol NO}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \times 18 \text{ g H}_2\text{O}}{1} = 6.75 \text{ g H}_2\text{O}$$

Percent Yield Problems And Answers

Heather Hattori, Richard H. Langley



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Table of Contents Percent Yield Problems And Answers

1. Understanding the eBook Percent Yield Problems And Answers
 - The Rise of Digital Reading Percent Yield Problems And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Percent Yield Problems And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Percent Yield Problems And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Yield Problems And Answers
 - Personalized Recommendations
 - Percent Yield Problems And Answers User Reviews and Ratings

- Percent Yield Problems And Answers and Bestseller Lists
- 5. Accessing Percent Yield Problems And Answers Free and Paid eBooks
 - Percent Yield Problems And Answers Public Domain eBooks
 - Percent Yield Problems And Answers eBook Subscription Services
 - Percent Yield Problems And Answers Budget-Friendly Options
- 6. Navigating Percent Yield Problems And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Percent Yield Problems And Answers Compatibility with Devices
 - Percent Yield Problems And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Percent Yield Problems And Answers
 - Highlighting and Note-Taking Percent Yield Problems And Answers
 - Interactive Elements Percent Yield Problems And Answers
- 8. Staying Engaged with Percent Yield Problems And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Percent Yield Problems And Answers
- 9. Balancing eBooks and Physical Books Percent Yield Problems And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Percent Yield Problems And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Percent Yield Problems And Answers
 - Setting Reading Goals Percent Yield Problems And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Percent Yield Problems And Answers
 - Fact-Checking eBook Content of Percent Yield Problems And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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