KEY

#### Percent, Actual, and Theoretical Yield

LiOH + KCl → LiCl + KOH

a) I began this reaction with 20 grams of fithium hydroxide. What is my theoretical yield of lithium chloride?

205 CiOH | mol CiOH | lad Cal | 42.39 CiCl = 35.405 CiCl)

b) I actually produced 6 grams of lithium chloride. What is my percent yield?

35. 40 CC × 100 = [16.9%

2) C<sub>3</sub>H<sub>8</sub> + 5 O<sub>2</sub> → 3 CO<sub>2</sub> + 4 H<sub>2</sub>O

RMCI + CWO + P CWOL + RM

a) If I start with 5 grams of C<sub>2</sub>H<sub>8</sub>, what is my theoretical yield of water?

B 5g C3Hg | and C3Hg | Had \$20 | 18.02 160 = 1 8.175 Hz 0 |

29 (00) 260 / W/ Reg 61.48 Re - 22.10, M. 0/

# **Percent Actual And Theoretical Yield Key**

Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George

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