

Periodic Table Basics

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

Step 2: Determine the number of protons, neutrons, and electrons in each element.

Step 3: Create a Bohr diagram for each element.

Step 4: Draw the Lewis Structure for each element.

The card for Boron (B) contains the following information:

- Atomic number: 5
- Symbol: B
- Name: Boron
- Atomic mass: 10.81
- Protons (P): 5
- Neutrons (N): 6
- Electrons (E): 5
- Bohr Diagram: A diagram showing a nucleus with 5 protons and 6 neutrons, and 5 electrons in two shells (2 in the first, 3 in the second).
- Lewis structure: A diagram showing the Boron atom with 5 valence electrons represented by dots around the symbol B.

Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

Green = Li & Na

Pink = O & S

Blue = Be & Mg

Purple = F & Cl

Orange = B & Al

Red = C & Si

Tan = N & P

Yellow = He, Ne, & Ar

Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.



Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.

Periodic Table Basics Answer Key

Kaplan Test Prep



Periodic Table Basics Answer Key:

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Basics in Nursing Research and Biostatistics Sreevani Rentala, 2018-10-30

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radiopharmaceutical chemistry an essential component of nuclear medicine and radiology However at this point interest in the field outpaces the academic and educational infrastructure needed to train radiopharmaceutical chemists For example the vast majority of texts that address radiopharmaceutical chemistry do so only peripherally focusing instead on nuclear chemistry i e nuclear reactions in reactors heavy element radiochemistry i e the decomposition of radioactive waste or solely on the clinical applications of radiopharmaceuticals e g the use of PET tracers in oncology This text fills that gap by focusing on the chemistry of radiopharmaceuticals with key coverage of how that knowledge translates to the development of diagnostic and therapeutic radiopharmaceuticals for the clinic The text is divided into three overarching sections First Principles Radiochemistry and Special Topics The first is a general overview covering fundamental and broad issues like The Production of Radionuclides and Basics of Radiochemistry The second section is the main focus of the book In this section each chapter s author will delve much deeper into the subject matter covering both well established and state of the art techniques in radiopharmaceutical chemistry This section will be divided according to radionuclide and will include chapters on radiolabeling methods using all of the common nuclides employed in radiopharmaceuticals including four chapters on the ubiquitously used fluorine 18 and a Best of the Rest chapter to cover emerging radionuclides Finally the third section of the book is dedicated to special topics with important information for radiochemists including Bioconjugation Methods Click Chemistry in Radiochemistry and Radiochemical Instrumentation This is an ideal educational guide for nuclear medicine physicians radiologists and radiopharmaceutical chemists as well as residents and trainees in all of these areas **ENC**

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