

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- a. O C Be Ne Same Period
 - b. Na Rb Fr H Same Group
 - c. Pb C Sn Si Same Group
 - d. Au W S Fr Ne Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- a. O C Be Ne Same Period
- b. Na Rb Fr H Same Group
- c. Pb C Sn Si Same Group
- d. Au W S Fr Ne Zn Challenge

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Barrett



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Trends in Teaching Experimentation in the Life Sciences* Nancy J. Pelaez, Stephanie M. Gardner, Trevor R. Anderson, 2022-05-11 This book is a guide for educators on how to develop and evaluate evidence based strategies for teaching biological experimentation to thereby improve existing and develop new curricula It unveils the flawed assumptions made at the classroom department and institutional level about what students are learning and what help they might need to develop competence in biological experimentation Specific case studies illustrate a comprehensive list of key scientific competencies that unpack what it means to be a competent experimental life scientist It includes explicit evidence based guidelines for educators regarding the teaching learning and assessment of biological research competencies The book also provides practical teacher guides and exemplars of assignments and assessments It contains a complete analysis of the variety of tools developed thus far to assess learning in this domain This book contributes to the growth of public understanding of biological issues including scientific literacy and the crucial importance of evidence based decision making around public policy It will be beneficial to life science instructors biology education researchers and science administrators who aim to improve teaching in life science departments Chapters 6 12 14 and 22 are available open access under a Creative Commons Attribution 4 0 International License via link [springer.com](https://www.springer.com)

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