



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Atoms share electrons to form a bond. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** One atom transfers electrons to another atom. Example: $NaCl$, MgO .
- Metallic Bond:** Atoms in a metal share their valence electrons in a "sea of electrons". Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonds

Covalent bonds are formed by sharing electrons between two non-metal atoms. Ionic bonds are formed by the transfer of electrons from a metal atom to a non-metal atom. Metallic bonds are formed by the sharing of valence electrons among many metal atoms.

Properties of Bonds:

- Covalent bonds are strong and directional.
- Ionic bonds are strong and non-directional.
- Metallic bonds are strong and non-directional.

Diagram: A periodic table of elements showing the general locations of elements that form covalent, ionic, and metallic bonds.



IONIC BONDING

Atoms combine to form ionic compounds by transferring electrons. The resulting compound has properties different from the individual atoms.

Diagram: A diagram showing the transfer of electrons from a metal atom to a non-metal atom, resulting in an ionic bond.

Properties of Ionic Bonds:

- Strong and non-directional.
- Form crystalline solids.
- High melting and boiling points.

Diagram: A diagram showing the structure of an ionic compound, with alternating positive and negative ions.

COVALENT BONDING

Atoms combine to form covalent compounds by sharing electrons. The resulting compound has properties different from the individual atoms.

Diagram: A diagram showing the sharing of electrons between two non-metal atoms, resulting in a covalent bond.

Properties of Covalent Bonds:

- Strong and directional.
- Form molecules or networks.
- Lower melting and boiling points than ionic compounds.

Diagram: A diagram showing the structure of a covalent compound, with atoms sharing electrons in pairs.

METALLIC BONDING

Atoms in a metal share their valence electrons in a "sea of electrons". The resulting metal has properties different from the individual atoms.

Diagram: A diagram showing the structure of a metal, with atoms in a lattice and a sea of delocalized electrons.

Properties of Metallic Bonds:

- Strong and non-directional.
- Form crystalline solids.
- High melting and boiling points.

GUIDED READING ACTIVITY

Reading Comprehension On Ionic And Covalent Bonds For Middle School

AN Whitehead



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Table of Contents Reading Comprehension On Ionic And Covalent Bonds For Middle School

1. Understanding the eBook Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - The Rise of Digital Reading Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Advantages of eBooks Over Traditional Books
2. Identifying Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Personalized Recommendations
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School User Reviews and Ratings
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School and Bestseller Lists

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5. Accessing Reading Comprehension On Ionic And Covalent Bonds For Middle School Free and Paid eBooks
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School Public Domain eBooks
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School eBook Subscription Services
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School Budget-Friendly Options
6. Navigating Reading Comprehension On Ionic And Covalent Bonds For Middle School eBook Formats
 - ePub, PDF, MOBI, and More
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School Compatibility with Devices
 - Reading Comprehension On Ionic And Covalent Bonds For Middle School Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Highlighting and Note-Taking Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Interactive Elements Reading Comprehension On Ionic And Covalent Bonds For Middle School
8. Staying Engaged with Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reading Comprehension On Ionic And Covalent Bonds For Middle School
9. Balancing eBooks and Physical Books Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reading Comprehension On Ionic And Covalent Bonds For Middle School
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Setting Reading Goals Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Fact-Checking eBook Content of Reading Comprehension On Ionic And Covalent Bonds For Middle School
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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