

# Answers: Oxidation Number

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1. N in  $\text{N}_2\text{O}_3$   $\text{N}^{3+}$

2. S in  $\text{H}_2\text{SO}_4$   $\text{S}^{6+}$

3. C  $\text{C}^0$

4. C in CO  $\text{C}^{2+}$

5. Na in NaCl  $\text{Na}^+$

6. H in  $\text{H}_2\text{O}$   $\text{H}^+$

7. Ba in  $\text{BaCl}_2$   $\text{Ba}^{2+}$

8. N in  $\text{NO}_2^-$   $\text{N}^{3+}$

9. S in  $\text{Al}_2\text{S}_3$   $\text{S}^{2-}$

10. S in  $\text{HSO}_4^-$   $\text{S}^{6+}$

11. Cl in  $\text{Fe}(\text{ClO}_2)_3$   $\text{Cl}^{3+}$

12. Fe in  $\text{Fe}(\text{ClO}_2)_3$   $\text{Fe}^{3+}$

13. N in  $\text{NO}_3^-$   $\text{N}^{5+}$

14. Cu<sup>2+</sup>  $\text{Cu}^{2+}$

15. Zn<sup>2+</sup>  $\text{Zn}^{2+}$

16. C in  $\text{CH}_4$   $\text{C}^{4+}$

17. Mn in  $\text{MnO}_2$   $\text{Mn}^{4+}$

18. S in  $\text{SO}_3^{2-}$   $\text{S}^{4+}$

19.  $\text{Mg}^{2+}$   $\text{Mg}^{2+}$

20.  $\text{Cl}^-$   $\text{Cl}^-$

21.  $\text{O}_2$   $\text{O}_2^0$

22.  $\text{P}_4$   $\text{P}_4^0$

23. Na in  $\text{Na}_2\text{S}$   $\text{Na}^+$

24. S in  $\text{H}_2\text{S}$   $\text{S}^{2-}$

25.  $\text{Ca}^{2+}$   $\text{Ca}^{2+}$

26. C in  $\text{CN}^-$   $\text{C}^{4+}$

27. H in  $\text{OH}^-$   $\text{H}^+$

28. Mn in  $\text{KMnO}_4$   $\text{Mn}^{7+}$

29. I in  $\text{Mg}(\text{IO}_3)_2$   $\text{I}^{5+}$

30. C in  $\text{C}_2\text{O}_4^{2-}$   $\text{C}^{2+}$

# Oxidation Numbers Section Review Answers Answers

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