

AP Physics – Photoelectric Effect

1. What energy is carried by a photon of electromagnetic radiation that has a frequency of 1.55×10^{17} Hz?

$$E = hf = 6.63 \times 10^{-34} \text{ J} \cdot \text{s} \left(1.55 \times 10^{17} \frac{1}{\text{s}} \right) = 10.3 \times 10^{-17} \text{ J} = \boxed{1.03 \times 10^{-16} \text{ J}}$$

2. A photon has a wavelength of 550 nm. How much energy does this represent in Joules?

$$v = f\lambda \quad \text{so} \quad c = \lambda f \quad f = \frac{c}{\lambda} \quad E = hf \quad E = \frac{hc}{\lambda}$$

$$E = \frac{(6.63 \times 10^{-34} \text{ J} \cdot \text{s}) (3.00 \times 10^8 \frac{\text{m}}{\text{s}})}{(550 \times 10^{-9} \text{ m})} = 0.036 \times 10^{-22} \text{ J} = \boxed{3.6 \times 10^{-24} \text{ J}}$$

3. What is the maximum kinetic energy of a photoelectron that has been liberated from a silver metal surface by a photon that has a frequency of 3.13×10^{15} Hz?

$$K_{Max} = hf - \phi = (4.14 \times 10^{-15} \text{ eV} \cdot \text{s}) \left(3.13 \times 10^{15} \frac{1}{\text{s}} \right) - 4.73 \text{ eV}$$

$$K_{Max} = \boxed{8.22 \text{ eV}}$$

4. What is the velocity of a photoelectron that has been liberated from a zinc metal surface by a photon that has a wavelength of 275 nm?

Consulting the table, we find that the work function for zinc is 4.31 eV. We can use this and the wavelength of the incident photons to find the kinetic energy of the ejected electrons. We can then solve for their velocity.

$$v = c = f\lambda \quad f = \frac{c}{\lambda} \quad E = hf = h \left(\frac{c}{\lambda} \right) = \frac{hc}{\lambda}$$

$$K_{Max} = \frac{hc}{\lambda} - \phi = \frac{1.24 \times 10^3 \text{ eV} \cdot \text{nm}}{275 \text{ nm}} - 4.31 \text{ eV} = 0.00451 \times 10^3 \text{ eV} - 4.31 \text{ eV}$$

$$K_{Max} = 4.51 \text{ eV} - 4.31 \text{ eV} = 0.200 \text{ eV}$$

$$K_{Max} = 0.200 \text{ eV} \left(\frac{1.60 \times 10^{-19} \text{ J}}{1 \text{ eV}} \right) = 0.32 \times 10^{-19} \text{ J} = \boxed{3.2 \times 10^{-20} \text{ J}}$$

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VM Jensen



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1. Understanding the eBook Pearson Photoelectric Effect Answer Key
 - The Rise of Digital Reading Pearson Photoelectric Effect Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Photoelectric Effect Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Photoelectric Effect Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Photoelectric Effect Answer Key
 - Personalized Recommendations
 - Pearson Photoelectric Effect Answer Key User Reviews and Ratings
 - Pearson Photoelectric Effect Answer Key and Bestseller Lists

5. Accessing Pearson Photoelectric Effect Answer Key Free and Paid eBooks
 - Pearson Photoelectric Effect Answer Key Public Domain eBooks
 - Pearson Photoelectric Effect Answer Key eBook Subscription Services
 - Pearson Photoelectric Effect Answer Key Budget-Friendly Options
6. Navigating Pearson Photoelectric Effect Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Photoelectric Effect Answer Key Compatibility with Devices
 - Pearson Photoelectric Effect Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Photoelectric Effect Answer Key
 - Highlighting and Note-Taking Pearson Photoelectric Effect Answer Key
 - Interactive Elements Pearson Photoelectric Effect Answer Key
8. Staying Engaged with Pearson Photoelectric Effect Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Photoelectric Effect Answer Key
9. Balancing eBooks and Physical Books Pearson Photoelectric Effect Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Photoelectric Effect Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Photoelectric Effect Answer Key
 - Setting Reading Goals Pearson Photoelectric Effect Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Photoelectric Effect Answer Key
 - Fact-Checking eBook Content of Pearson Photoelectric Effect Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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