

A gloved hand is pouring a bright orange liquid from a graduated cylinder into a petri dish. The background is a blurred laboratory setting with a Bunsen burner visible.

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**Yijin Wang**



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web gabrielle m 9 years ago i m working on chemical reactions double and single replacement on flvs now my first question for this video is how do you have 4 aluminum atoms when it says  $2\text{Al}$  subscript 2 do you just add the 2 and the subscript 2 same thing goes with the  $\text{O}_3$  i am so lost

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web balancing equations and simple stoichiometry answers balance the following equations  $\text{N}_2$   $\text{F}_2$   $\text{NF}_3$   $\text{C}_6\text{H}_{10}\text{O}_2$   $\text{CO}_2$   $\text{H}_2\text{O}$   $3\text{HBr}$   $\text{KHCO}_3$   $\text{H}_2\text{O}$   $\text{KBr}$   $\text{CO}_2$   $4\text{GaBr}_3$   $\text{Na}_2\text{SO}_3$   $\text{Ga}_2\text{SO}_3$   $3\text{S}$   $5\text{SNO}$   $\text{NF}_3$   $\text{N}_2\text{O}$   $\text{CO}_2$

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web google classroom balance the following chemical equation  $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$  note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem 7 4 1 x x y y theta  $\theta$  pi  $\pi$  8 5 2 0 9 6

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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