Section: Calculating Quantities in Reactions

Complete each statement below by writing the correct term or phrase.

- When solving stoichiometric problems, you must ______ the equation first.
- Balanced equations give the ______ numbers of moles of substances.
- in chemical equations provide mole ratios that can be used as conversion factors.
- The conversion factor for converting between mass and moles is the of the substance.
- In making calculations involving ______, you must convert volume to mass.
- To convert from volume to mass, you can use the ______ of the substance as the conversion factor.
- 8. When calculating the number of particles, you can use _____ as the conversion factor.

In the blanks at the left, write the letter of the choice that best completes the statement or answers the question. Consider the following problem when answering:

What mass of sulfuric acid is required to neutralize 2.65 g of potassium hydroxide? The products of the reaction are potassium sulfate and water.

- 9. What should you do first after reading the problem carefully?
 - a. Estimate the answer.
 - b. Calculate the molar mass of sulfuric acid.
 - c. Write a balanced chemical equation.
 - d. Convert all masses to moles.
- 10. What should you do before setting up the problem?
 - a. Determine the densities.
 - b. Calculate molar masses.
 - c. Convert all masses to moles.
 - d. Estimate the answer.

Skills Worksheet Concept Review Answers Section What Is Radioactivity

Huangqi Zhang

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