# 18 (pt. 2) • Precipitation Reactions

#### STUDY QUESTIONS

 Predict, on the basis of the solubility rules, which of the following salts are soluble and which are insoluble. For those that are "insoluble" look up their solubility products.

a. AgI c. BaSO<sub>4</sub> e. NiCO<sub>3</sub> b. Na<sub>3</sub>PO<sub>4</sub> d. (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> f. Cu(OH)<sub>2</sub>

2. Write solubility product expressions for the following salts:

a. PbSO<sub>4</sub> c. CuS

b. Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>
 d. CaF<sub>2</sub>

- If the molar concentration of lead bromide, PbBr<sub>2</sub>, in an aqueous solution is 1.6 x 10<sup>-6</sup> M, what is [Pb<sup>2+</sup>] and [Br\*]?
- If the molar solubility of silver iodide is 1.22 x 10<sup>-8</sup> M, what is the solubility product for AgI?
- What is the molar solubility of cadmium sulfide, CdS, if its K<sub>sp</sub> = 3.6 x 10<sup>-29</sup>?
- K<sub>sp</sub> of strontium fluoride, SrF<sub>2</sub>, is 2.5x 10<sup>-9</sup>. What is the [Sr<sup>2+</sup>] and [F<sup>+</sup>] in a saturated solution of SrF<sub>2</sub>?
  What is the molar solubility of SrF<sub>2</sub>?
- 7. What is the [Mg<sup>2+</sup>] in a saturated solution of magnesium fluoride, MgF<sub>2</sub> if its solubility product constant is 6.4 x 10<sup>-9</sup>? What is the [Mg<sup>2+</sup>] if the solution also contains 0.30 M sodium fluoride?

# **Pearson Science Work Answers Precipitation**

J Ma

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