

**13****STATES OF MATTER****Practice Problems****SECTION 13.1 THE NATURE OF GASES**

1. Explain why there is no gas pressure inside a vacuum.

*Gas pressure is the result of collisions between rapidly moving particles in a gas and an object. Because there are no particles of matter in a vacuum, there can be no collisions or pressure.*

2. How would the reading on a barometer change if you were to take one on a trip from Los Angeles to Lake Tahoe, which is at a much higher altitude?

*Setting aside fluctuations due to changes in the weather, you would notice that the pressure reading on the barometer would decrease as you climbed in altitude.*

3. The height of a column of mercury in a barometer is 754.3 mm. What is the atmospheric pressure in atm? In kPa?

*0.9925 atm*

*100.5 kPa*

4. How does the average kinetic energy of the helium atoms in a balloon change as the helium gas is heated from  $-100.0^{\circ}\text{C}$  to  $73^{\circ}\text{C}$ ?

*The average kinetic energy of the particles of a substance is directly proportional to the Kelvin Temperature.*

$$-100.0^{\circ}\text{C} + 273 = 173\text{ K}$$

$$73^{\circ}\text{C} + 273 = 346\text{ K}$$

*Because the Kelvin temperature increases by a factor of two, the average kinetic energy increases by a factor of two.*

# Pearson States Of Matter Practice Problems

**Robert D. Postman**



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## **Table of Contents Pearson States Of Matter Practice Problems**

1. Understanding the eBook Pearson States Of Matter Practice Problems
  - The Rise of Digital Reading Pearson States Of Matter Practice Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson States Of Matter Practice Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson States Of Matter Practice Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson States Of Matter Practice Problems
  - Personalized Recommendations
  - Pearson States Of Matter Practice Problems User Reviews and Ratings
  - Pearson States Of Matter Practice Problems and Bestseller Lists
5. Accessing Pearson States Of Matter Practice Problems Free and Paid eBooks
  - Pearson States Of Matter Practice Problems Public Domain eBooks
  - Pearson States Of Matter Practice Problems eBook Subscription Services
  - Pearson States Of Matter Practice Problems Budget-Friendly Options

6. Navigating Pearson States Of Matter Practice Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - Pearson States Of Matter Practice Problems Compatibility with Devices
  - Pearson States Of Matter Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson States Of Matter Practice Problems
  - Highlighting and Note-Taking Pearson States Of Matter Practice Problems
  - Interactive Elements Pearson States Of Matter Practice Problems
8. Staying Engaged with Pearson States Of Matter Practice Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson States Of Matter Practice Problems
9. Balancing eBooks and Physical Books Pearson States Of Matter Practice Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson States Of Matter Practice Problems
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson States Of Matter Practice Problems
  - Setting Reading Goals Pearson States Of Matter Practice Problems
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson States Of Matter Practice Problems
  - Fact-Checking eBook Content of Pearson States Of Matter Practice Problems
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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