

Assessment

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Understanding Concepts

- What is a stable electron configuration?
- What does each dot in an electron dot diagram represent?
- What process changes atoms into ions?
- What keeps the ions in their fixed positions within a crystal lattice?
- What are subscripts used for in chemical formulas?
- Explain why a melted ionic compound is a good conductor of electric current, but a solid ionic compound is a poor conductor of electric current.
- What distinguishes single, double, and triple covalent bonds?
- Explain why the covalent bonds in molecules of O_2 and N_2 are always nonpolar.
- Explain why, in a covalent bond between oxygen and hydrogen, the hydrogen atom has a partial positive charge and the oxygen atom has a partial negative charge.
- What is the name of the binary compound formed from potassium and iodine?
- Write the formulas for the compounds called copper(I) chloride and copper(II) chloride.
- Name the compounds represented by the space-filling models labeled A, B, and C.



- In general, what determines the strength of metallic bonds?
- What properties of copper and tin change when these metals are mixed together to form bronze?
- What advantage of magnesium is retained in magnesium alloys? What disadvantage is reduced?

Chemical Bonds 187

Assessment



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Reviewing Content

- | | | |
|-------|------|------|
| 1. b | 2. c | 3. a |
| 4. d | 5. c | 6. b |
| 7. d | 8. b | 9. c |
| 10. a | | |

Understanding Concepts

- In a stable electron configuration, the highest occupied energy level is filled with electrons.
- Each dot represents a valence electron.
- The transfer of electrons.
- Attractions between neighboring cations and anions keep the ions in fixed positions within the lattice.
- A subscript is used to show the number of atoms of an element in a molecule or the ratio of ions in a crystal lattice.
- When an ionic compound melts, ions can move away from their fixed locations in the crystal lattice.
- Two atoms share two electrons in a single bond, four in a double bond, and six in a triple bond.
- The covalent bonds in molecules of elements are always nonpolar because the atoms have the same attraction for electrons.
- The oxygen atom has a greater attraction for electrons than the hydrogen atom does.
- Potassium iodide
- CuCl and CuCl_2
- A is sulfur trioxide, B is carbon monoxide, and C is nitrogen dioxide
- In general, the more valence electrons a metal has, the stronger the metal bonds are.
- A mixture of copper and tin is harder and stronger than either metal in its pure form.
- The advantage that is retained is that magnesium is a lightweight metal. The disadvantage that is reduced is that magnesium is a soft metal.

Index

- atoms
11–16, 26, 30, 33
8, 17–19, 27, 34–35, 39–40, 42
20–22, 28–29, 31–32, 36–38
23–25, 41

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The Peirce Edition Project



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1. Understanding the eBook Pearson Physical Science Chapter 4 Review Answers
 - The Rise of Digital Reading Pearson Physical Science Chapter 4 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Physical Science Chapter 4 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Physical Science Chapter 4 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Physical Science Chapter 4 Review Answers
 - Personalized Recommendations
 - Pearson Physical Science Chapter 4 Review Answers User Reviews and Ratings
 - Pearson Physical Science Chapter 4 Review Answers and Bestseller Lists
5. Accessing Pearson Physical Science Chapter 4 Review Answers Free and Paid eBooks
 - Pearson Physical Science Chapter 4 Review Answers Public Domain eBooks
 - Pearson Physical Science Chapter 4 Review Answers eBook Subscription Services
 - Pearson Physical Science Chapter 4 Review Answers Budget-Friendly Options
6. Navigating Pearson Physical Science Chapter 4 Review Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Pearson Physical Science Chapter 4 Review Answers Compatibility with Devices
 - Pearson Physical Science Chapter 4 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Physical Science Chapter 4 Review Answers
 - Highlighting and Note-Taking Pearson Physical Science Chapter 4 Review Answers
 - Interactive Elements Pearson Physical Science Chapter 4 Review Answers
 8. Staying Engaged with Pearson Physical Science Chapter 4 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Physical Science Chapter 4 Review Answers
 9. Balancing eBooks and Physical Books Pearson Physical Science Chapter 4 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Physical Science Chapter 4 Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pearson Physical Science Chapter 4 Review Answers
 - Setting Reading Goals Pearson Physical Science Chapter 4 Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pearson Physical Science Chapter 4 Review Answers
 - Fact-Checking eBook Content of Pearson Physical Science Chapter 4 Review Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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