

# NAME: \_\_\_\_\_ EQUILIBRIUM PRACTICE

## Equilibrium Practice

Name: \_\_\_\_\_

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1. The reaction below has reached a state of equilibrium. Write the equilibrium constant expression for this reaction.
2. Consider the following chemical equation:  $2\text{H}_2\text{O}(g) \rightleftharpoons 2\text{H}_2(g) + \text{O}_2(g)$ . In equilibrium, the concentration of  $\text{H}_2$  is 0.01 M, the concentration of  $\text{O}_2$  is 0.005 M, and the concentration of  $\text{H}_2\text{O}$  is 0.10 M. Calculate the equilibrium constant,  $K$ .
3. At a certain pressure, 10 g of each of the gases, chlorine and hydrogen, were placed in a 1.00 L container. At equilibrium, the concentration of  $\text{H}_2$  was 0.01 M. Calculate the equilibrium constant,  $K$ .
4. A 1.00 L flask contains 0.100 mol of  $\text{H}_2$  and 0.100 mol of  $\text{I}_2$ . At equilibrium, the concentration of  $\text{H}_2$  is 0.050 M. Calculate the equilibrium constant,  $K$ .
5. Consider the following reaction:  $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$ . In the initial concentration of  $\text{PCl}_5$  is 0.100 M. At equilibrium, the concentration of  $\text{Cl}_2$  is 0.020 M. Calculate the equilibrium constant,  $K$ .

## Skill Practice 45 Equilibrium Practice Answers

**Patricia A. Potter, Anne G.  
Perry, Patricia A. Stockert, Amy Hall**

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