

Periodic Trends: Ionization Energy Worksheet

- What is the ionization energy of an element?
- Predict the following trends you expect to see in ionization energy:
 - As you move across a period? Why?
 - As you move down a group? Why?
- What factors influence the ease of removing an electron from an element?
- Choose an orbital where an electron would experience the highest Z_{eff} , effective nuclear charge (least shielded), and the highest first ionization energy. Circle the correct answer.

| | | | | |
|---------|---------|---------|--------|--------|
| Na (3s) | Mg (3s) | Al (3p) | P (3p) | S (3p) |
|---------|---------|---------|--------|--------|
- Match the following electron configuration with the appropriate ionization energy.

| | |
|--|-----------------|
| a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ | i) 1356 kJ/mol |
| b) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ | ii) 595 kJ/mol |
| c) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$ | iii) 409 kJ/mol |
- Which one of the following has the largest first ionization energy? Circle the correct answer.

| | | | |
|----|---|----|----|
| Na | K | Li | Cs |
|----|---|----|----|
- The first ionization energies values (kJ/mol) for sequential elements in period 3 are shown below. Assign elements to these values based on the periodic trends.

| | | | | |
|-----|-----|-----|------|-----|
| 740 | 578 | 786 | 1012 | 999 |
|-----|-----|-----|------|-----|
- Choose the element with the greatest first ionization energy. Circle the correct answer.

| | |
|-----------------------|-------------------------|
| a) Carbon or aluminum | b) Calcium or strontium |
| c) Helium or lithium | d) Chlorine or argon |

Skill Practice 17 Ionization Answers

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