

**Section 18.4 Neutralization***In your textbook, read about neutralization and titration.*

For each item in Column A, write the letter of the matching item in Column B.

**Column A**

- \_\_\_\_\_ 1. A chemical dye that changes color based on the pH of a solution
- \_\_\_\_\_ 2. A method for using a neutralization reaction to determine the concentration of a solution
- \_\_\_\_\_ 3. A reaction in which an acid and a base react to produce a salt and water
- \_\_\_\_\_ 4. A solution of known concentration
- \_\_\_\_\_ 5. An ionic product of an acid-base reaction
- \_\_\_\_\_ 6. The point in a titration in which an indicator changes color
- \_\_\_\_\_ 7. The stoichiometric point of a titration

**Column B**

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. salt
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

Acid	Base	Salt formula	Salt name
8. HCl	KOH	KCl	potassium chloride
9. $\text{H}_2\text{SO}_4$	$\text{Mg}(\text{OH})_2$		
10. $\text{H}_3\text{PO}_4$	NaOH		
11. $\text{HNO}_3$	$\text{Fe}(\text{OH})_3$		
12. $\text{H}_3\text{PO}_4$	$\text{Ca}(\text{OH})_2$		

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

**Sequence of Steps**

- \_\_\_\_\_ 13. Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- \_\_\_\_\_ 14. Fill a buret with the standard solution.
- \_\_\_\_\_ 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

# Review Stoichiometry Section 1 And 2 Answers

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## **Table of Contents Review Stoichiometry Section 1 And 2 Answers**

1. Understanding the eBook Review Stoichiometry Section 1 And 2 Answers
  - The Rise of Digital Reading Review Stoichiometry Section 1 And 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Review Stoichiometry Section 1 And 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Review Stoichiometry Section 1 And 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Review Stoichiometry Section 1 And 2 Answers
  - Personalized Recommendations
  - Review Stoichiometry Section 1 And 2 Answers User Reviews and Ratings
  - Review Stoichiometry Section 1 And 2 Answers and Bestseller Lists
5. Accessing Review Stoichiometry Section 1 And 2 Answers Free and Paid eBooks
  - Review Stoichiometry Section 1 And 2 Answers Public Domain eBooks
  - Review Stoichiometry Section 1 And 2 Answers eBook Subscription Services
  - Review Stoichiometry Section 1 And 2 Answers Budget-Friendly Options

6. Navigating Review Stoichiometry Section 1 And 2 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Review Stoichiometry Section 1 And 2 Answers Compatibility with Devices
  - Review Stoichiometry Section 1 And 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Review Stoichiometry Section 1 And 2 Answers
  - Highlighting and Note-Taking Review Stoichiometry Section 1 And 2 Answers
  - Interactive Elements Review Stoichiometry Section 1 And 2 Answers
8. Staying Engaged with Review Stoichiometry Section 1 And 2 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Review Stoichiometry Section 1 And 2 Answers
9. Balancing eBooks and Physical Books Review Stoichiometry Section 1 And 2 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Review Stoichiometry Section 1 And 2 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Review Stoichiometry Section 1 And 2 Answers
  - Setting Reading Goals Review Stoichiometry Section 1 And 2 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review Stoichiometry Section 1 And 2 Answers
  - Fact-Checking eBook Content of Review Stoichiometry Section 1 And 2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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