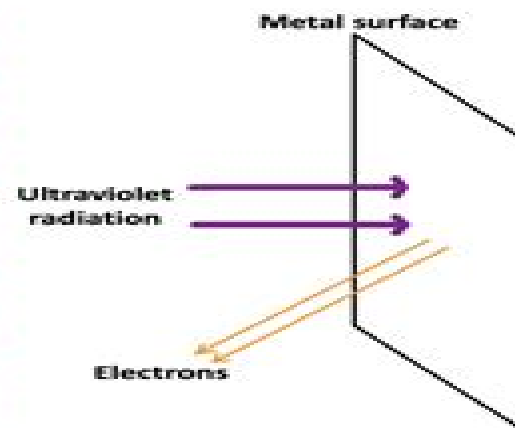


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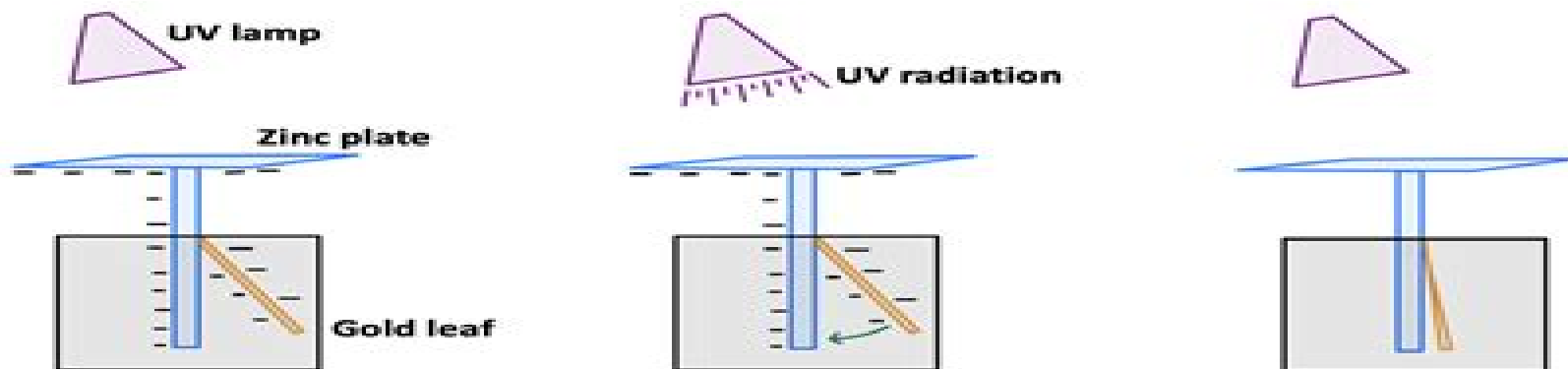
1) Explain, with a labelled diagram, the photoelectric effect.

In the photoelectric effect, the electromagnetic wave is thought to exist as quanta (discrete packets) of energy, called **photons**. When an EM wave is directed at a metal surface, the **photons are absorbed by the electrons**, causing them to be emitted from the surface. This will only happen if the **photon has greater energy than the work function energy of the metal**. The kinetic energy of the electrons depends on the energy of the photon. The maximum KE of an electron emitted is equal to the energy of the photon minus the work function of the metal. The number of electrons emitted depends on light intensity, as a higher light intensity means more photons of light and therefore more electrons emitted.



2) Explain, with a labelled diagram, the gold-leaf electroscope experiment.

A zinc plate is attached to the electroscope's top and given a **negative charge**, causing both the metal and the gold leaf inside the box to become negatively charged. The **repulsion** between the negatively charged metal and the gold leaf causes the **leaf to rise**. Then, **ultraviolet (UV) light** is directed onto the zinc plate. The energy from the light triggers the **photoelectric effect**, leading to the **loss of electrons** from the zinc plate. As the negatively charged entities (zinc plate, metal, and gold leaf) lose their charge, the **repulsion diminishes**, causing the **gold leaf to fall back down**.



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