

# BALANCING CHEMICAL EQUATIONS

Name Grove

Rewrite and balance the equations below.

- $N_2 + H_2 \rightarrow NH_3$   $N_2 + 3H_2 \rightarrow 2NH_3$
- $KClO_3 \rightarrow KCl + O_2$   $2KClO_3 \rightarrow 2KCl + 3O_2$
- $NaCl + F_2 \rightarrow NaF + Cl_2$   $2NaCl + F_2 \rightarrow 2NaF + Cl_2$
- $H_2 + O_2 \rightarrow H_2O$   $2H_2 + O_2 \rightarrow 2H_2O$
- $AgNO_3 + MgCl_2 \rightarrow AgCl + Mg(NO_3)_2$   $2AgNO_3 + MgCl_2 \rightarrow 2AgCl + Mg(NO_3)_2$
- $AlBr_3 + K_2SO_4 \rightarrow KBr + Al_2(SO_4)_3$   $2AlBr_3 + 3K_2SO_4 \rightarrow 6KBr + Al_2(SO_4)_3$
- $CH_4 + O_2 \rightarrow CO_2 + H_2O$   $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$
- $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$   $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$
- $C_8H_{18} + O_2 \rightarrow CO_2 + H_2O$   $C_8H_{18} + \frac{25}{2}O_2 \rightarrow 8CO_2 + 9H_2O$
- $FeCl_3 + NaOH \rightarrow Fe(OH)_3 + NaCl$   $FeCl_3 + 3NaOH \rightarrow Fe(OH)_3 + 3NaCl$
- $P + O_2 \rightarrow P_2O_5$   $4P + 5O_2 \rightarrow 2P_2O_5$
- $Na + H_2O \rightarrow NaOH + H_2$   $2Na + 2H_2O \rightarrow 2NaOH + H_2$
- $Ag_2O \rightarrow Ag + O_2$   $2Ag_2O \rightarrow 4Ag + O_2$
- $S_8 + O_2 \rightarrow SO_3$   $S_8 + 12O_2 \rightarrow 8SO_3$
- $CO_2 + H_2O \rightarrow C_6H_{12}O_6 + O_2$   $6CO_2 + 6H_2O \rightarrow C_6H_{12}O_6 + 6O_2$
- $K + MgBr_2 \rightarrow KBr + Mg$   $2K + MgBr_2 \rightarrow 2KBr + Mg$
- $HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$   $2HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$

# Physical Science If8767 Worksheet Answers Chemical Equations

**Georg Job, Regina Rüffler**



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Chemistry Equations and Answers Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students

**Chemical Equations: Chemistry Essentials Practice Workbook with 200+ Reactions to Balance** Jake Mariska,2021-05-08 A chemical equation is something you will encounter every day in chemistry It s a written representation using numbers and symbols of the process that occurs during a chemical reaction Master the art of balancing chemical reactions through examples and practice 10 examples are fully solved step by step with explanations to serve as a guide Over 200 chemical equations provide ample practice Exercises start out easy and grow progressively more challenging and involved Answers to every problem are tabulated at the back of the book A chapter of pre balancing exercises helps develop essential counting skills Opening chapter reviews pertinent concepts and ideas

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**Chemistry, Problem-Solving Worktext** James E. Brady, Frederick Senese, 2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo.

**Fireworks** One of the ingredients is phosphorus, a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms. The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light.

**Carbon** The element is found in several forms, including the familiar diamond and another recently discovered sooty substance that consists of soccer ball-shaped molecules, often referred to as buckyballs. Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth.

**Acetylene torch** Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen, a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$ .

**Hot air balloon** The air that helps it rise is heated by the combustion of molecules of propane, each composed of three carbon and eight hydrogen atoms.

**Stormy weather** The evaporation of water serves to store energy provided by the sun. Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet.

**Physical Chemistry for the Life Sciences Solutions Manual** Maria Bohorquez, 2005-07-26 The Solutions Manual is a powerful study aid that contains the complete answers to all the exercises in the text These worked out solutions guide you through each step and help you refine your problem solving skills Used in conjunction with the text the Solutions Manual is one of the best ways to develop a fuller appreciation of chemical principles It can also be used to review material identify problem areas where more study is needed and test yourself before an exam Book jacket

**Student Solutions Manual to Accompany a Conceptual Introduction to Chemistry** Rich Bauer, James Birk, Pamela S. Marks, Pamela Marks, 2006-02 The Student Solutions Manual contains detailed solutions and explanations for all odd numbered problems in the text

*Holt Science Spectrum Physical Science Chapter 7 Resource File: Chemical Reactions* Holt Rinehart & Winston, 2008-01-01

**Balancing Chemical Equations**, 2009 Inquiries in Science Chemistry Series Balancing Chemical Equations Teacher's Guide

*The Ultimate Chemical Equations Handbook* George R. Hague, Jane D. Smith, 2001

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