

Write the number of each item of silver in the following compounds. (2 points each)

33.)  $\text{Mg}(\text{NO}_3)_2$      Mg     N     O

- a. 4-4-4     b. 1-4-4  
c. 1-1-4     d. 1-1-1

34.)  $\text{Cu}_2(\text{PO}_3)_4$      Cu     P     O

- a. 2-2-3     b. 2-2-6  
c. 2-1-6     d. 2-2-6

35.)  $(\text{NH}_4)_2\text{SO}_4$      N     S     O

- a. 2-4-2-7     b. 1-4-2-7  
c. 4-4-2-7     d. 4-4-2-7

Draw a cooling curve phase diagram.

36.)  $\text{H}_2\text{O}$

- a. Solid     b. Gas  
c. Liquid     d. Melting  
e. Freezing

37.)  $\text{CO}_2$

- a. Solid     b. Gas  
c. Liquid     d. Melting  
e. Freezing

38.)  $\text{CO}_2$

- a. Solid     b. Gas  
c. Liquid     d. Melting  
e. Freezing

39.) Which element has the highest molar mass?

- a.  $\text{H}_2$      b.  $\text{O}_2$   
c.  $\text{Cl}_2$      d.  $\text{N}_2$   
e.  $\text{Br}_2$

40.) Which element has the lowest molar mass?

- a.  $\text{H}_2$      b.  $\text{O}_2$   
c.  $\text{Cl}_2$      d.  $\text{N}_2$   
e.  $\text{Br}_2$



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**Arshad Iqbal**



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