

## Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed, you should predict the products, then balance. Further questions just have the reactants listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of redox reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't worry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know forms ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

1. Combustion:  $\text{C}_2\text{H}_6 + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
2. Combustion:  $2 \text{C}_4\text{H}_{10} + 11 \text{O}_2 \rightarrow 8 \text{CO}_2 + 6 \text{H}_2\text{O}$
3. Combustion:  $\text{C}_5\text{H}_{12}\text{O}_2 + 7 \text{O}_2 \rightarrow 6 \text{CO}_2 + 5 \text{H}_2\text{O}$
4. Synthesis:  $\text{Mg} + \text{I}_2 \rightarrow \text{MgI}_2$
5. Double displacement:  $\text{CuCl}_2 + \text{H}_2\text{S} \rightarrow \text{CuS} + 2 \text{HCl}$
6. Double displacement:  $\text{NaOH} + \text{HClO}_4 \rightarrow \text{NaClO}_4 + \text{H}_2\text{O}$
7. Decomposition:  $\text{ZnCO}_3 + \text{heat} \rightarrow \text{ZnO} + \text{CO}_2$
8. Single replacement:  $2 \text{HCl} + \text{Zn} \rightarrow \text{ZnCl}_2 + \text{H}_2$
9. single  $2 \text{Na} + \text{MgCl}_2 \rightarrow 2 \text{NaCl} + \text{Mg}$
10. double  $\text{CaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2 \text{KCl}$
11. synthesis  $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$
12. double  $3 \text{BaCl}_2 + 2 \text{K}_3\text{PO}_4 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6 \text{KCl}$
13. double  $\text{H}_2\text{SO}_4 + 2 \text{KOH} \rightarrow \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
14. synthesis  $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
15. double  $\text{Pb}(\text{NO}_3)_2 + 2 \text{KOH} \rightarrow \text{Pb}(\text{OH})_2 + 2 \text{KNO}_3$
16. double  $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow 2 \text{HCl} + \text{BaSO}_4$
17. single  $\text{Ca} + 2 \text{AgCl} \rightarrow \text{CaCl}_2 + 2 \text{Ag}$
18. double  $\text{H}_3\text{PO}_4 + \text{FeBr}_3 \rightarrow 3 \text{HBr} + \text{FePO}_4$
19. synthesis  $6 \text{Li} + \text{N}_2 \rightarrow 2 \text{Li}_3\text{N}$
20. double  $2 \text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + 2 \text{H}_2\text{O}$

# Writing Equations Predicting Reactions Answer Key

**Richard Bailey**



## **Writing Equations Predicting Reactions Answer Key:**

## Embracing the Song of Appearance: An Psychological Symphony within **Writing Equations Predicting Reactions Answer Key**

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### **Table of Contents Writing Equations Predicting Reactions Answer Key**

1. Understanding the eBook Writing Equations Predicting Reactions Answer Key
  - The Rise of Digital Reading Writing Equations Predicting Reactions Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Writing Equations Predicting Reactions Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Writing Equations Predicting Reactions Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Writing Equations Predicting Reactions Answer Key

- Personalized Recommendations
  - Writing Equations Predicting Reactions Answer Key User Reviews and Ratings
  - Writing Equations Predicting Reactions Answer Key and Bestseller Lists
5. Accessing Writing Equations Predicting Reactions Answer Key Free and Paid eBooks
    - Writing Equations Predicting Reactions Answer Key Public Domain eBooks
    - Writing Equations Predicting Reactions Answer Key eBook Subscription Services
    - Writing Equations Predicting Reactions Answer Key Budget-Friendly Options
  6. Navigating Writing Equations Predicting Reactions Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Writing Equations Predicting Reactions Answer Key Compatibility with Devices
    - Writing Equations Predicting Reactions Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Writing Equations Predicting Reactions Answer Key
    - Highlighting and Note-Taking Writing Equations Predicting Reactions Answer Key
    - Interactive Elements Writing Equations Predicting Reactions Answer Key
  8. Staying Engaged with Writing Equations Predicting Reactions Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Writing Equations Predicting Reactions Answer Key
  9. Balancing eBooks and Physical Books Writing Equations Predicting Reactions Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Writing Equations Predicting Reactions Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Writing Equations Predicting Reactions Answer Key
    - Setting Reading Goals Writing Equations Predicting Reactions Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Writing Equations Predicting Reactions Answer Key

- Fact-Checking eBook Content of Writing Equations Predicting Reactions Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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