CHEMISTRY 11 – WORD EQUATIONS AND TYPES OF CHEMICAL REACTIONS WORKSHEET

- Translate the following word equations into chemical symbols and balance the resulting equations.
 Do not include the phases.
 - (a) Potassium + water → potassium hydroxide + hydrogen

(b) Nitrogen + hydrogen → ammonia

- In each of the following write a balanced chemical equation, including the phases.
 - (a) Liquid phosphoric acid reacts with aqueous barium hydroxide to give water and a precipitate of barium phosphate

(b) Solid aluminium oxide and aqueous sulphuric acid produce water and aqueous aluminium sulphate

(c) Nitrogen trifluoride gas and hydrogen gas react to form nitrogen gas and gaseous hydrogen fluoride

(d) Powdered sodium carbonate and aqueous hydrobromic acid react to form carbon dioxide gas, aqueous sodium bromide and water

Classify the following reactions.

Reaction	Туре
2 C ₂ H ₂ + 5 O ₂ → 4 CO ₂ + 2 H ₂ O	Combustion
$Mg + CuSO_4 \rightarrow MgSO_4 + Cu$	Single Replacement
4 Na + O₂ → 2 Na₂O	Synthesis
$Sn(OH)_4 + 4 HBr \rightarrow 4 H_2O + SnBr_4$	Neutralization
2 Fe(NO ₃) ₃ + 3 MgS → Fe ₂ S ₃ + 3 Mg(NO ₃) ₂	Double Replacement
2 N ₂ O → 2 N ₂ + O ₂	Decomposition

Writing Chemical Equations Rxn Answer Key

Kenneth W. Raymond

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