

Writing Complete Equations Practice - Key

For each of the following problems, write complete chemical equations to describe the chemical process taking place. Important note: There are a few physical processes on this sheet – remember, you can't write an equation for a physical process!

- 1) When lithium hydroxide pellets are added to a solution of sulfuric acid, lithium sulfate and water are formed.



- 2) When dirty water is boiled for purification purposes, the temperature is brought up to 100° C for 15 minutes.

No equation is needed, as boiling is a physical process.

- 3) If a copper coil is placed into a solution of silver nitrate, silver crystals form on the surface of the copper. Additionally, highly soluble copper (I) nitrate is generated.



- 4) When crystalline $\text{C}_6\text{H}_{12}\text{O}_6$ is burned in oxygen, carbon dioxide and water vapor are formed.



- 5) When a chunk of palladium metal is ground into a very fine powder and heated to drive off any atmospheric moisture, the resulting powder is an excellent catalyst for chemical reactions.

Both grinding and heating are physical processes. Even if the atmospheric moisture is mentioned, boiling is still a physical process. No equation is needed.

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. **The Practice of Chemistry Study Guide & Solutions Manual** Pamela

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