

# Empirical and Molecular Formulas Worksheet

Key

## Objectives:

- be able to calculate empirical and molecular formulas

## Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} = 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} = 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1 \end{array}$$



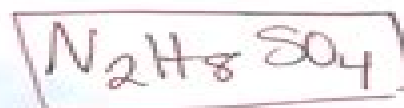
- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l} \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \\ \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5 \end{array}$$



- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l} \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2 \\ \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8 \\ \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1 \\ \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4 \end{array}$$



## Molecular Formula

- 4) Empirical formula of a substance is  $\text{CH}_2\text{O}$ . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) = 30$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	$\text{Au}_2\text{O}_3$
3	$\text{NH}_4\text{SO}_4$
4	$\text{C}_2\text{H}_8\text{S}_4\text{O}_4$
5	$\text{C}_6\text{H}_{12}\text{O}_6$

# Writing Empirical And Molecular Formulas Answers

**Sasha Hu**



## **Writing Empirical And Molecular Formulas Answers:**

## Reviewing **Writing Empirical And Molecular Formulas Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Writing Empirical And Molecular Formulas Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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### **Table of Contents Writing Empirical And Molecular Formulas Answers**

1. Understanding the eBook Writing Empirical And Molecular Formulas Answers
  - The Rise of Digital Reading Writing Empirical And Molecular Formulas Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Writing Empirical And Molecular Formulas Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Writing Empirical And Molecular Formulas Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Writing Empirical And Molecular Formulas Answers
  - Personalized Recommendations
  - Writing Empirical And Molecular Formulas Answers User Reviews and Ratings
  - Writing Empirical And Molecular Formulas Answers and Bestseller Lists

5. Accessing Writing Empirical And Molecular Formulas Answers Free and Paid eBooks
  - Writing Empirical And Molecular Formulas Answers Public Domain eBooks
  - Writing Empirical And Molecular Formulas Answers eBook Subscription Services
  - Writing Empirical And Molecular Formulas Answers Budget-Friendly Options
6. Navigating Writing Empirical And Molecular Formulas Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Writing Empirical And Molecular Formulas Answers Compatibility with Devices
  - Writing Empirical And Molecular Formulas Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Writing Empirical And Molecular Formulas Answers
  - Highlighting and Note-Taking Writing Empirical And Molecular Formulas Answers
  - Interactive Elements Writing Empirical And Molecular Formulas Answers
8. Staying Engaged with Writing Empirical And Molecular Formulas Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Writing Empirical And Molecular Formulas Answers
9. Balancing eBooks and Physical Books Writing Empirical And Molecular Formulas Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Writing Empirical And Molecular Formulas Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Writing Empirical And Molecular Formulas Answers
  - Setting Reading Goals Writing Empirical And Molecular Formulas Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Writing Empirical And Molecular Formulas Answers
  - Fact-Checking eBook Content of Writing Empirical And Molecular Formulas Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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