

Electron Configuration-Unit 4

Orbital Diagrams and Electron Configuration:

Use the patterns within the periodic table to draw orbital diagrams and write the electron configuration for the following terms:

Symbol	#e ⁻	Orbital diagrams and Electron Configuration
O	8	$1s^2 2s^2 2p^4$
Ge	32	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^2$ or [Ar] $4s^2 3d^{10} 4p^2$
N	7	$1s^2 2s^2 2p^3$
K	19	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ [Ar] $4s^1$
Cr	24	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^5$ [Ar] $4s^1 3d^5$
Se	34	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$ or [Ar] $4s^2 3d^{10} 4p^4$
Hg	80	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$ or [Xe] $6s^2 4f^{14} 5d^{10}$
Ti	22	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^3$ or [Ar] $4s^1 3d^3$
Pb	82	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$ or [Xe] $6s^2 4f^{14} 5d^{10} 6p^2$
Ca	20	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ or [Ar] $4s^2$
F	9	$1s^2 2s^2 2p^5$
Pd	46	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^9$ [Kr] $5s^1 4d^9$
Br	35	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$ [Ar] $4s^2 3d^{10} 4p^5$
Ar	18	$1s^2 2s^2 2p^6 3s^2 3p^6$
Te	52	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6$ [Kr] $5s^2 4d^{10} 5p^6$
Zn	30	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10}$ [Ar] $4s^2 3d^{10}$
Bi	83	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^3$ [Xe] $6s^2 4f^{14} 5d^{10} 6p^3$
Cs	55	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^1$ [Xe] $6s^1$

Ws Orbital Notation Answers

Guillaume Favre

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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