

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Calculate:** All limiting reagent must limit everything that comes along. **Remember:** Limiting reagent always comes up first when figuring out.

Chemists will determine the limiting reagent by calculating the amount of product that each reactant will form. The reactant that makes the least product is the limiting reagent.



The reactant required with the smallest amount for the product is the limiting reagent.

2H ₂ (g)	20.0 mol product
O ₂ (g)	10.0 mol product
2H ₂ O(l)	20.0 mol product
H ₂ O	10.0 mol product

1. 20.0 g of each reactant will be used. 1 mol is also 10.0 g. 20.0 g of each reactant means 20.0 g of H₂ and 20.0 g of O₂. **Remember:** Limiting reagent will limit everything that comes along. **Remember:** Limiting reagent always comes up first when figuring out.



2. **Remember:** Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Limiting reagent always comes up first when figuring out.

3. 20.0 g of each reactant will be used. 1 mol is also 10.0 g. 20.0 g of each reactant means 20.0 g of H₂ and 20.0 g of O₂. **Remember:** Limiting reagent will limit everything that comes along. **Remember:** Limiting reagent always comes up first when figuring out.



4. 20.0 g of each reactant will be used. 1 mol is also 10.0 g. 20.0 g of each reactant means 20.0 g of H₂ and 20.0 g of O₂. **Remember:** Limiting reagent will limit everything that comes along. **Remember:** Limiting reagent always comes up first when figuring out.



Percent Yield And Limiting Reactants Quiz

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