

Periodic Table Basics

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

Step 2: Determine the number of protons, neutrons, and electrons in each element.

Step 3: Create a Bohr diagram for each element.

Step 4: Draw the Lewis Structure for each element.

The sample card for Boron (B) contains the following information:

- Atomic number: 5
- Symbol: B
- Name: Boron
- Atomic mass: 10.81
- Protons (P): 5
- Neutrons (N): 6
- Electrons (E): 5
- Bohr Diagram: A diagram showing a nucleus with 5 protons and 6 neutrons, and 5 electrons arranged in two shells (2 in the first shell, 3 in the second shell).
- Lewis structure: A diagram showing the Boron atom with 5 valence electrons represented by dots around the symbol B.

Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

Green = Li & Na

Pink = O & S

Blue = Be & Mg

Purple = F & Cl

Orange = B & Al

Red = C & Si

Tan = N & P

Yellow = He, Ne, & Ar

Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.



Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.

Periodic Table Basics Answers

William C. Robertson, Michael S. Kralik



Periodic Table Basics Answers:

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 *Study Guide to Accompany Basics for Chemistry* is an 18 chapter text designed to be used with *Basics for Chemistry* textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter, and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds, chemical formulas, and the mole, chemical reactions, calculations based on equations, gases, and the properties of a liquid. The remaining chapters examine the solutions, acids, bases, salts, oxidation-reduction reactions, electrochemistry, chemical kinetics, and equilibrium, and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students.

The Periodic Table Becky Ham, 2009 A short guide to one of the most celebrated diagrams in the history of science, a tool for understanding the basic building blocks of the universe. **The Complete Idiot's Guide to the GED**

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Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble,1994 Chemistry Workbook For Dummies Chris Hren,Peter J.

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