

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O C Be Ne Same Period
 - Na Rb Fr H Same Group
 - Pb C Sn Si Same Group
 - Au W S Fr Ne Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O C Be Ne Same Period
- Na Rb Fr H Same Group
- Pb C Sn Si Same Group
- Au W S Fr Ne Zn Challenge

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H Kauffman

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