



Trotter 410t By Cybex Manual

Yan Bai



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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction time the total pressure $p_a = x_a p_t$ mole fraction $x_a = \frac{\text{moles } a}{\text{moles total}}$ $\sum x_i = 1$ mole fraction of substance a is often denoted by x_a and

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mass C_6H_6 6 12 6 1 g mol 1 78 g mol 1 hence C_6H_6 s number of moles $\frac{30.78 \text{ mol}}{0.3846 \text{ mol}}$ molar mass of carbon tet CCl_4 12 4 355 154 g mol 1

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the o 2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of o 2 are mixed in a flask at room temperature