

EXPERIMENT 8: Activity Series (Single Displacement Reactions)

PURPOSE

- Reactions of metals with acids and salt solutions
- Determine the activity of metals
- Write a balanced molecular equation, complete ionic equation, and net ionic equation for single displacement reaction.

Materials:

Equipment: 2 well-plates (Spot plates) or small test tubes

Stirring rod or toothpicks

Chemicals and solutions: a tray containing the metals Mg, Pb, Cu, Zn, Fe, 1.0 M HCl and 0.25 M nitrate salt solutions of each of these metals such as:

$\text{Mg}(\text{NO}_3)_2$

$\text{Pb}(\text{NO}_3)_2$

$\text{Cu}(\text{NO}_3)_2$

$\text{Zn}(\text{NO}_3)_2$

$\text{Fe}(\text{NO}_3)_3$

Note - Before you begin this lab, you should be familiar with types of reactions found in chapter 4 of the text. Wear appropriate clothing, shoes and safety glasses for this lab.



Stirring rod



Spot plate

INTRODUCTION

One of the most important characteristics of a metal is its activity (reactivity). The activity of a metal is its ability to react with nonmetals. In such reaction a metal typically loses electrons to the non-metal to form cations and the product is an ionic compound.

Different metals have different activities. Some metals are very active such as lithium, sodium, potassium, and cesium; some are slowly active such as tin, lead, and iron; and some are not active at all (inactive) such as copper, silver, and mercury. Inactive metals are those which resist oxidation (reaction with oxygen in air). These are often called the noble metals.

Single Replacement Lab Key

Guillaume Favre

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