Reaction Kinetics and Equilibrium

9.5 Energy in Chemical Reactions

- system & surroundings
- enthalpy
- heat of reaction(Δ enthalpy) & calculations
- endothermic and exothermic processes

I Can Statements:

1. I can understand the reason that chemical reactions take place in terms of relative stability and collision theory. (13.1, 9.5)

A. I can explain the involvement of energy transfer in an exothermic reaction.

 I can use a graph to diagram and identify parts of an exothermic reaction.

- a. reactants
- b. products
- c. activation energy
- d. activated complex
- e. enthalpy change (∆H).

B. I can explain the involvement of energy transfer in an endothermic reaction.

> I can use a graph to diagram and identify parts of an exothermic reaction.

- a, reactants
- b. products
- c. activation energy
- d. activated complex
- e. enthalpy change (4H)

I can identify factors that affect the reaction rate and explain their effects on reaction rate. (13.1)

- A. Temperature
- B. Concentration of Reactants
- C. Pressure
- D. Surface Area
- C. Presence of a Catalyst

 I can show a catalyst's effect on activation energy given a reaction graph.

I can use the concept of reversible reactions to explain chemical equilibrium.
(13.2, 13.5)

Reaction Energy And Reaction Kinetics Study Guide

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