



## What you will do

### Self-Test 2.1

Try to check how much you have learned in this lesson by answering the following questions.

I. Write the equilibrium expression for the following reactions

1.  $\text{Ni}_{(s)} + 4\text{CO}_{(g)} \rightleftharpoons \text{Ni(CO)}_{4(g)}$
2.  $\text{H}_2\text{O}_{(g)} \rightleftharpoons \text{H}_2\text{O}_{(l)}$
3.  $\text{NiO}_{(s)} + \text{CO}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{Ni}_{(s)}$
4.  $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$
5.  $3\text{H}_{2(g)} + \text{N}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$

II. Consider the exothermic reaction



Write  $\rightarrow$  if the reaction will go forward,  $\leftarrow$  for reverse and  $\rightleftharpoons$  for no change in the direction of the reaction.

- |       |                             |
|-------|-----------------------------|
| _____ | 1. Adding oxygen            |
| _____ | 2. Removing sulfur dioxide  |
| _____ | 3. Removing sulfur trioxide |
| _____ | 4. Adding helium gas        |
| _____ | 5. Adding catalyst          |



Key to answers on page 15.

## Lesson 3. Applications of Chemical Equilibrium

In this lesson, we will study some of the wonderful applications of chemical equilibrium. The principles of chemical equilibrium are evident from the production of industrial chemicals, the acid-balance in lakes and streams, to the maintenance of our blood pH.

### Haber Process: Ammonia synthesis

We have encountered the synthesis of ammonia from previous lessons. This time, we will take a closer look at this reaction. The complete equation is written as



# Reaction Rates And Equilibrium Lesson Check Answers

**Jeff Lark**



## **Reaction Rates And Equilibrium Lesson Check Answers:**

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