

- 2.77 (a) Li^{+} , alkali metals always have a +1 charge in ionic compounds.
 (b) S^{2-} .
 (c) I^{-} , halogens have a -1 charge in ionic compounds.
 (d) N^{3-} .
 (e) Al^{3+} , aluminum always has a +3 charge in ionic compounds.
 (f) Ca^{2+} , alkali metals always have a +1 charge in ionic compounds.
 (g) Mg^{2+} , alkaline earth metals always have a +2 charge in ionic compounds.

2.78 The symbol ^{23}Na provides more information than Na . The mass number plus the chemical symbol identifies a specific isotope of Na (sodium) while combining the atomic number with the chemical symbol tells you nothing new. Can other isotopes of sodium have different atomic numbers?

2.79 The binary Group 7A element acids are: HF, hydrofluoric acid; HCl, hydrochloric acid; HBr, hydrobromic acid; HI, hydroiodic acid. Oxacids containing Group 7A elements (using the specific examples for chlorine) are: HClO_4 , perchloric acid; HClO_3 , chloric acid; HClO_2 , chlorous acid; HClO , hypochlorous acid.

Examples of oxacids containing other Group A-block elements are: H_2BO_3 , boric acid (Group 3A); H_2CO_3 , carbonic acid (Group 4A); HNO_3 , nitric acid and H_3PO_4 , phosphoric acid (Group 5A); and H_2SO_4 , sulfuric acid (Group 6A). Hydrofluoric acid, HF , is an example of a binary Group 6A acid while HCN , hydrocyanic acid, contains both a Group 4A and 5A element.

2.80 Mercury (Hg) and bromine (Br)

2.81 (a)	Isotope	^{214}Po	^{210}Po	^{210}At	^{210}Bi	^{210}Po
	No. Protons	84	84	85	83	84
	No. Neutrons	130	126	125	127	126
(b)	neutron/proton ratio	1.55	1.50	1.47	1.53	1.51

The neutron/proton ratio increases with increasing atomic number.

2.82 H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , I_2 , Ne , Ar , Kr , Xe , Rn

2.83 Cu, Ag, and Au are fairly chemically unreactive. This makes them specially suitable for making coins and jewelry, that you want to last a very long time.

2.84 They do not have a strong tendency to form compounds. Helium, neon, and argon are chemically inert.

2.85 Magnesium and calcium are also alkaline earth metals. You should expect the charge of the metal to be the same (+2). MgO and BaO .

2.86 All isotopes of radium are radioactive. It is a radioactive decay product of uranium-238. Radium itself does not occur naturally on Earth.

- 2.87 (a) Berkelium (Berkley, CA); Europium (Europe); Francium (France); Scandium (Scandinavia);
 Ytterbium (Ytterby, Sweden); Yttrium (Ytterby, Sweden).
 (b) Einsteinium (Albert Einstein); Fermium (Enrico Fermi); Curium (Marie and Pierre Curie);
 Mendelevium (Dmitri Mendeleev); Lawrencium (Ernest Lawrence).
 (c) Arsenic, Cesium, Chlorine, Chromium, Iodine.

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Ensheng Dong



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Table of Contents Raymond Chang 10th Edition Solution Manual

1. Understanding the eBook Raymond Chang 10th Edition Solution Manual
 - The Rise of Digital Reading Raymond Chang 10th Edition Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Raymond Chang 10th Edition Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Raymond Chang 10th Edition Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Raymond Chang 10th Edition Solution Manual
 - Personalized Recommendations
 - Raymond Chang 10th Edition Solution Manual User Reviews and Ratings
 - Raymond Chang 10th Edition Solution Manual and Bestseller Lists

5. Accessing Raymond Chang 10th Edition Solution Manual Free and Paid eBooks
 - Raymond Chang 10th Edition Solution Manual Public Domain eBooks
 - Raymond Chang 10th Edition Solution Manual eBook Subscription Services
 - Raymond Chang 10th Edition Solution Manual Budget-Friendly Options
6. Navigating Raymond Chang 10th Edition Solution Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Raymond Chang 10th Edition Solution Manual Compatibility with Devices
 - Raymond Chang 10th Edition Solution Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Raymond Chang 10th Edition Solution Manual
 - Highlighting and Note-Taking Raymond Chang 10th Edition Solution Manual
 - Interactive Elements Raymond Chang 10th Edition Solution Manual
8. Staying Engaged with Raymond Chang 10th Edition Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Raymond Chang 10th Edition Solution Manual
9. Balancing eBooks and Physical Books Raymond Chang 10th Edition Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Raymond Chang 10th Edition Solution Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Raymond Chang 10th Edition Solution Manual
 - Setting Reading Goals Raymond Chang 10th Edition Solution Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Raymond Chang 10th Edition Solution Manual
 - Fact-Checking eBook Content of Raymond Chang 10th Edition Solution Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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