

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

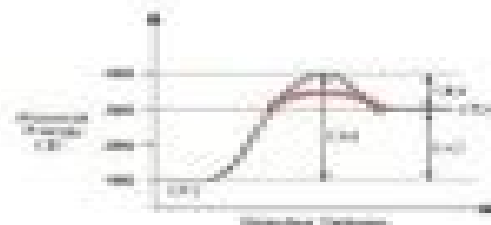
- 9-1. Describe heat, temperature, and heat flow in terms of the motion of molecules (or atoms). (C30, 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (C30, 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat (e.g. $m \cdot c \cdot \Delta T$) & latent heat of phase change (e.g. $m \cdot \Delta H$). (C30, 7d)
 9-4. Know the rate of reactions is affected by the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (C30, 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (C30, 8a*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (C30, 8e)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (C30, 8f)

Vocabulary

- | | | | |
|----------------|-------------------|--|----------------------------|
| • Condensation | • Deposition | • Endotherm | • Heat of reaction |
| • Freezing | • Heat | • Exotherm | • Le Chatelier's principle |
| • Vaporization | • Temperature | • Specific heat | • Equilibrium |
| • Melting | • Thermochemistry | • $q = m \cdot c \cdot \Delta T$ | • Activation Energy |
| • Boiling | • Exothermic | • $q = m \cdot \Delta H_{\text{vaporation}}$ | • Catalyst |
| • Sublimation | • Endothermic | • $q = m \cdot \Delta H_{\text{fusion}}$ | |

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
Must collide with each other, Must collide with enough energy, & in the correct orientation
- What is the importance of the activation energy in a chemical reaction?
Amount of energy needed for a reaction to occur
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:
 $2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
 - Increase temp: **Increases, molecules collide more**
 - Increase HCl: **Increases, more collisions because there are more reactants to collide**
 - Increase pressure: **Hard to say, wouldn't affect the reaction directly because none of them are gases. One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call... (that one won't be on the test)**
 - Remove MgCl_2 : **Increases, less products means less chance for them to interfere with reactant collisions**
 - Add H_2 : **Decreases, more products means they interfere more with reactants and the reactants would collide less**
- Identify the following terms with their correct letter from the graph:
 - Reactants- **J**
 - Products- **F**
 - Activation energy (number/ letter & amount)-
A (100-75) kcal
 - Is this reaction exothermic or endothermic?
Endothermic
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It **lowers the activation energy (see on graph). Speeds it up by lowering the activation energy.**



Reaction Rates And Equilibrium Worksheet Answers

Chapter 18

**Thobeka Pete, Turret Correspondence
College (Johannesburg)**



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