

OPINION WRITING RUBRIC

Area of Writing	 3	 2	 1	SCORE
Topic	<i>Fully introduces the topic or specifically names the book they are writing about</i>	<i>Vaguely introduces the topic or alludes to (but does not name) the book they are writing about</i>	<i>Does not introduce the topic or name the book they are writing about</i>	
Opinion & Reason	<i>States both an opinion and supplies a reason for the opinion</i>	<i>States an opinion or a reason for the opinion</i>	<i>Does not state an opinion or a reason for the opinion</i>	
Closure	<i>Provides a sense of closure related to the topic</i>	<i>Provides a sense of closure that is unrelated</i>	<i>Does not provide a sense of closure</i>	
Comments:				TOTAL
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Adapted From Mary Jo Jordan's Opinion Writing Rubric

Rubric For Common Core Opinion Letter

VM Jensen



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effect of impact shock and stresses are higher than rcc structures 2 longevity of prestressed structure is greater than rcc
structure because the reinforcement stays unaffected from outer agencies 3

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concrete but also high tensile steel wires concrete used for prestressed work should have cube strength of 35 n mm² for
post tensioned system and 45n m m² for pretensioned system

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web the prestressed concrete is much more usable in the construction of bridges railway sleepers and dams the dead load of the structure will be reduced by using prestressed concrete and it indirectly reduces the usage of steel and project cost disadvantages of prestressed concrete it is economical for large span structures only ample space

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economy as by adopting prestressing we can reduce the of the er for higher load smaller c s can be achieved a large span can be achieved without increasing the member size for the same loading condition

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jan 3 2023 a net ionic equation is a chemical equation that shows only the species participating in a chemical reaction and omits the spectator ions spectator ions are ions that

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds cl aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4^+ \text{aq} + 2\text{F}^- \text{aq}$ because two NH_4^+ aq

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