

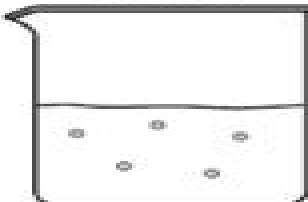
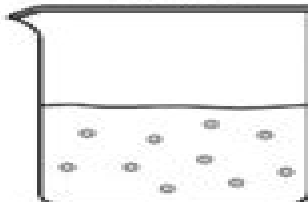
Saturated and Unsaturated Solutions

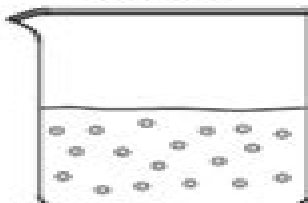
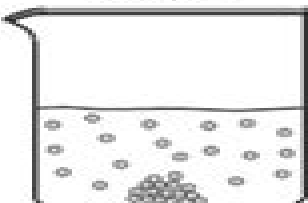
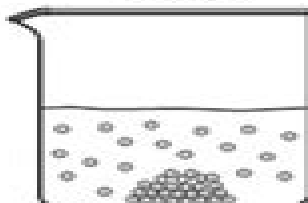
Is there a limit to the amount of solute that will dissolve in a solvent?

Why?

We use solutions every day. People who wear contact lenses use "lens solution" to rinse their contacts and keep them wet. Athletes who consume sports drinks after exercising benefit from the electrolytes in those solutions. This activity will explore whether or not there is a limit to how much of one substance can dissolve in another.

Model 1 – Saturated and Unsaturated Solutions

- All beakers contain 10.0 g of water. - All beakers are kept at 20 °C. - All solutions are stirred for 2 hours. - Solute is the same substance in all beakers.	Unsaturated Solutions	
	Beaker A  1.0 g of solute added	Beaker B  2.0 g of solute added
	Number of dissolved particles	5
	Number of solid particles	0

Saturated Solutions		
Beaker C  3.6 g of solute added	Beaker D  7.0 g of solute added	Beaker E  9.0 g of solute added
		18
	17	

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Peipei Pang



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virtual lab circuit basics phet contribution - Nov 27 2022

web mar 1 2018 virtual lab circuit basics description students start with basic circuit concepts building a simple circuit

recording the current and testing the equation $v = ir$ then they start building series and parallel circuits and develop the rules

for current and potential difference for each type