

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV=nRT$ ", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 kPa / 1 atm$ to get
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of $87^{\circ}C$, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of $56^{\circ}C$, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of $67^{\circ}C$, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

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David M. Hata



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Fundamentals of Chemical Engineering Thermodynamics Themis Matsoukas, 2013 *Fundamentals of Chemical Engineering Thermodynamics* is the clearest and most well organized introduction to thermodynamics theory and calculations for all chemical engineering undergraduates This brand new text makes thermodynamics far easier to teach and learn Drawing on his award winning courses at Penn State Dr Themis Matsoukas organizes the text for more effective learning focuses on why as well as how offers imagery that helps students conceptualize the equations and illuminates thermodynamics with relevant examples from within and beyond the chemical engineering discipline Matsoukas presents solved problems in every chapter ranging from basic calculations to realistic safety and environmental applications

Elementary Principles of Chemical Processes Richard M. Felder, Ronald W. Rousseau, Lisa G. Bullard, 2020-08-11 This best selling text prepares students to formulate and solve material and energy balances in chemical process systems and lays the foundation for subsequent courses in chemical engineering The text provides a realistic informative and positive introduction to the practice of chemical engineering *Applied Petroleum Reservoir Engineering* Ronald E. Terry, J. Brandon Rogers, Benjamin Cole Craft, 2015 This book presents many real field examples demonstrating the use of material balance and history matching to predict reservoir performance For the first time this edition uses Microsoft Excel with VBA as its calculation tool making calculations far easier and more intuitive for today's readers Beginning with an introduction of key terms detailed coverage of the material balance approach and progressing through the principles of fluid flow water influx and advanced recovery techniques this book will be an asset to students without prior exposure to petroleum engineering with this text updated to reflect modern industrial practice **The Software Encyclopedia**, 1988 Chemistry

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Principles and Calculations in Process Technology T. David Griffith, 2015-09-02 A Practical Guide to Physical and Chemical Principles and Calculations for Today's Process Control Operators In Basic Principles and Calculations in Process Technology author T David Griffith walks process technologists through the basic principles that govern their operations helping them collaborate with chemical engineers to improve both safety and productivity He shows process operators how to go beyond memorizing rules and formulas to understand the underlying science and physical laws so they can accurately interpret anomalies and respond appropriately when exact rules or calculation methods don't exist Using simple algebra and non technical analogies Griffith explains each idea and technique without calculus He introduces each topic by explaining why it matters to process technologists and offers numerous examples that show how key principles are applied and calculations are performed For end of chapter problems he provides the solutions in plain English discussions of how and why they work Chapter appendixes provide more advanced information for further exploration Basic Principles and Calculations in Process Technology is an indispensable practical resource for every process technologist who wants to know what the numbers mean so they can control their systems and processes more efficiently safely and reliably T David Griffith received his B S in chemical engineering from The University of Texas at Austin and his Ph D from the University of Wisconsin Madison then top ranked in the discipline After working in research on enhanced oil recovery EOR he cofounded a small chemical company and later in his career he developed a record setting Electronic Data Interchange EDI software package He currently instructs in the hydrocarbon processing industry Coverage includes Preparing to solve problems by carefully organizing them and establishing consistent sets of measures Calculating areas and volumes including complex objects and interpolation Understanding Boyle's Law Charles's Law and the Ideal Gas Law Predicting the behavior of gases under extreme conditions Applying thermodynamic laws to calculate work and changes in gas enthalpy and to recognize operational problems Explaining phase equilibria for distillation and fractionalization Estimating chemical reaction speed to optimize control Balancing material or energy as they cross system boundaries Using material balance calculations to confirm quality control and prevent major problems Calculating energy balances and using them to troubleshoot poor throughput Understanding fluid flow including shear viscosity laminar and turbulent flows vectors and tensors Characterizing the operation of devices that transport heat energy for heating or cooling Analyzing mass transfer in separation processes for materials purification

A Learning Program for Chemistry Peter Demmin, 1993-11

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Table of Contents Pearson Prentice Hall Answer Key Ideal Gases

1. Understanding the eBook Pearson Prentice Hall Answer Key Ideal Gases
 - The Rise of Digital Reading Pearson Prentice Hall Answer Key Ideal Gases
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Prentice Hall Answer Key Ideal Gases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Prentice Hall Answer Key Ideal Gases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Prentice Hall Answer Key Ideal Gases
 - Personalized Recommendations
 - Pearson Prentice Hall Answer Key Ideal Gases User Reviews and Ratings
 - Pearson Prentice Hall Answer Key Ideal Gases and Bestseller Lists
5. Accessing Pearson Prentice Hall Answer Key Ideal Gases Free and Paid eBooks
 - Pearson Prentice Hall Answer Key Ideal Gases Public Domain eBooks
 - Pearson Prentice Hall Answer Key Ideal Gases eBook Subscription Services
 - Pearson Prentice Hall Answer Key Ideal Gases Budget-Friendly Options
6. Navigating Pearson Prentice Hall Answer Key Ideal Gases eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Prentice Hall Answer Key Ideal Gases Compatibility with Devices
 - Pearson Prentice Hall Answer Key Ideal Gases Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Prentice Hall Answer Key Ideal Gases
 - Highlighting and Note-Taking Pearson Prentice Hall Answer Key Ideal Gases
 - Interactive Elements Pearson Prentice Hall Answer Key Ideal Gases
8. Staying Engaged with Pearson Prentice Hall Answer Key Ideal Gases

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Prentice Hall Answer Key Ideal Gases
- 9. Balancing eBooks and Physical Books Pearson Prentice Hall Answer Key Ideal Gases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Prentice Hall Answer Key Ideal Gases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Prentice Hall Answer Key Ideal Gases
 - Setting Reading Goals Pearson Prentice Hall Answer Key Ideal Gases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Prentice Hall Answer Key Ideal Gases
 - Fact-Checking eBook Content of Pearson Prentice Hall Answer Key Ideal Gases
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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