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# 13 STATES OF MATTER

### Practice Problems SECTION 13.1 THE NATURE OF GASES

1. Explain why there is no gas pressure inside a vacuum

Gas pressure is the result of collisions between rapidly moving particles in a gas and an object. Because there are no particles of matter in a vacuum, there can be no collisions or pressure.

2. How would the reading on a becometer change if you were to take one on a trip from Los Angeles to Lake Tahoe, which is at a much higher altitude?

Setting aside fluctuations due to changes in the weather, you would notice that the pressure reading on the barometer would decrease as you climbed in altitude.

 The height of a column of mercury in a becometer is 7543 mm. What is the atmospheric pressure in atm? In kPs?

0.9925 attm

JOO SkiPla

4. How does the everage kinetic energy of the helium atoms in a balloon change as the helium gas is heated from -100 0°C to 73°C?

The average kinetic energy of the particles of a substance is directly proportional to the Kelvin Temperature.

-100.0 \*C + 273 = 173 K

73 °C + 273 = 346 K

Because the Kelvin temperature increases by a factor of two, the average kinetic energy increases by a factor of two.

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