Section 1 Review page 118

1. Describe Democritus's atomic theory.

Everything in the universe is made up of atoms.

- 2. Summarize the main ideas of Dalton's theory.
 - Elements are made up of tiny unique particles called atoms.
 - b. Atoms cannot be divided.
 - c. Atoms of the same element are identical.
 - d. Atoms of different elements can join to form molecules.
- Explain why Dalton's theory was more successful tan Democritus's theory.
 Dalton had experimental evidence to support his theory, while Democritus did not.
- Compare Thomson's atomic model with Rutherford's atomic model.
 Thomson's model- electrons were embedded in a sphere of positive charge Rutherford's model- positive charge is concentrated in the nucleus and electrons surrounded the nucleus.
- 5. Analyzing elements
 - a. How did the cathode-ray tube experiment lead to the conclusion that atoms contain electrons? The air was removed from the tube and the beam of negative particles was observed. This suggests that atoms are made up of smaller units with negative charges.
 - b. How did the gold-foil experiment lead to the conclusion that the atom has a nucleus? the positive particles were deflected at large angles and therefore this suggests that a positive charge is concentrated in the center of the atom. (nucleus)
- Making Inferences: Does the term indivisible still describe the atom? Explain. No, the atom is not indivisible. It is made up of smaller parts such as the electron.

Section 2 review page 127

- List the charge, mass, and location of each of the three subatomic particles found in atoms. The proton is positive, the neutron is neutral and electrons are negative. Protons and neutrons are found in the nucleus of an atom. The mass of a proton 1850 times greater than that of an electron.
- 2. Explain how you can use an atom's mass number and atomic number to determine the number of protons, electrons, and neutrons to determine the number of protons, electrons, and neutrons in the atom. The atomic number equals the number of protons. The number of electrons equals the number of protons. To find the number of neutrons subtract the atomic number from the mass number.
- Identify the subatomic particle used to define an element, and explain why
 this particle is used. The proton defines the element's number because the each
 element always has the same number of protons, while the number of neutrons
 and electrons can vary.
- Explain why the masses of atoms of the same element may differ. Because they can have different number of neutrons. For example Hydrogen (H) H, H₁Deterium, H₂ Tritium. These are called isotopes.

Pearson Physical Science Chapter 4 Review

Marek Bryjak, Nalan Kabay, Bernabe L. Rivas, Jochen Bundschuh

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Physics for College Students Ansel Alphonso Knowlton,1928 Measurement and Evaluation in Physical Activity *Applications* Phillip A. Bishop, 2017-06-30 This book takes a learner oriented approach as it strives to make complex material understandable and usable By understanding the underlying principles of measurement and evaluation readers will then be able to apply those principles and concepts in a variety of physical activity and health related settings Practical exercises and applications demonstrate the usefulness of measurement and evaluation reinforce key points and make readers active participants in their own education The book is divided into three parts Part One introduces the measurement process showing readers the relevance of measurement and evaluation to their personal and professional lives and including examples and statistics related to such concepts as validity reliability and objectivity. The two chapters in Part Two further help readers understand numbers and assist those who need to use more advanced statistical calculations Part Three presents measurement and evaluation applications in various settings such as measuring physical fitness measuring exercise physical activity and health measuring in competitive sports and coaching measuring and evaluating knowledge and assigning grades and measuring in research Throughout discussions and examples show the relevance and application of measurement and evaluation in various professions including physical therapy athletic training fitness wellness management exercise and sport psychology exercise science coaching and physical education *Innovative Materials and Methods for* Water Treatment Marek Bryjak, Nalan Kabay, Bernabe L. Rivas, Jochen Bundschuh, 2016-02-17 Due to increasing demand for potable and irrigation water water suppliers have to use alternative resources. They either have to regenerate wastewater or deal with contaminated surface water This book brings together the experiences of various experts in preparing of innovative materials that are selective for arsenic and chromium removal and in The Scientific Monthly James McKeen **Objectivity: A Very Short Introduction** Stephen Gaukroger, 2012-05-24 Objectivity is both an essential Cattell,1923 and elusive philosophical concept This Very Short Introduction explores the theoretical and practical problems raised by objectivity and also deals with the way in which particular understandings of objectivity impinge on social research science Telegraphic Journal and Monthly Illustrated Review of Electrical Science ,1924 and art Electrochemical *Energy Storage* Reinhart Job, 2020-09-21 Starting from physical and electrochemical foundations this textbook explains working principles of energy storage devices After a history of galvanic cells different types of primary secondary and flow cells as well as fuel cells and supercapacitors are covered An emphasis lies on the general setup and mechanisms behind those devices to enable easy understanding for students from all technical and natural science disciplines Materials for Semiconductor Functions E. G. Bylander,1971 Life Span Human Development Carol K. Sigelman,Linda De George, Kimberley Cunial, Elizabeth A. Rider, 2018-09-01 The third edition of Life Span Human Development helps students gain a deeper understanding of the many interacting forces affecting development from infancy childhood adolescence and

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