## Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO<sub>3</sub> (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	2 KClO <sub>3</sub>	→ 2 KCI +	3 02
Before:	1.2	•	0
Change:	-1.2	+1.2	+1.8
After:	0	1.2	1.8

1.2 moles KClO<sub>2</sub> needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equations	Nig +	2 HCl →	MgCl <sub>1</sub> +	HI a
Before:	0.100	7	0	0
Change:	-0.100	-0.200 +	0.100 +	0.100
Afters		7	0.100	0.100

0.200 moles HCI needed

3. Ethane, C<sub>2</sub>H<sub>6</sub> reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	2 C2H6	+ 7 O <sub>2</sub>	→ 4 CO;	+ 6.H <sub>2</sub> O
Before:	2.20	76.76	0	
Changes	-2.20	- 7.70	+ 4,40	+ 6.60
After:	0	3.5	4.40	6.60

O<sub>2</sub> needed: 7.70 mole x 32.0 g = 246 g O<sub>2</sub> required 1 mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:

Ni(NO<sub>3</sub>)<sub>2</sub> + 2 NaOH --> Ni(OH)<sub>2</sub> + 2 NaNO<sub>3</sub>

0.73 g x 1 mole = .0040 mole 183 g

Equation:	Ni(NO <sub>3</sub> ):	2 NaOH	→ Ni(OH)2	+ 2 NaNO3
Before:	0.0040	XS	0	0
Change:	- 0.0040	- 0.0080	+0.0040	+0.0080
After:	0	NS	0.0040	0.0080

NaNO, produced: 0.0080 mole x 85.0 g = 0.68 g I mole

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