

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste

Feel slippery

Contains hydroxide ions

Turns litmus blue

Reacts with metals to form H_2

Turns litmus red

Bitter taste

Contains hydrogen ions

Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an H^+ ion in solution	Donates an OH^- ion in solution
Bronsted-Lowry	Donates an H^+ ion	Accepts an H^+ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the $[H^+]$ is greater than the $[OH^-]$, the solution is acidic basic neutral
6. If the $[OH^-]$ is greater than the $[H^+]$, the solution is acidic basic neutral
7. If the $[H^+]$ is equal to the $[OH^-]$, the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3 / 84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH / 1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3 / 63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2 / 2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2 / 1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

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 - The Rise of Digital Reading Review Acids And Bases Modern Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Acids And Bases Modern Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Acids And Bases Modern Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Acids And Bases Modern Chemistry Answers
 - Personalized Recommendations
 - Review Acids And Bases Modern Chemistry Answers User Reviews and Ratings
 - Review Acids And Bases Modern Chemistry Answers and Bestseller Lists
5. Accessing Review Acids And Bases Modern Chemistry Answers Free and Paid eBooks
 - Review Acids And Bases Modern Chemistry Answers Public Domain eBooks
 - Review Acids And Bases Modern Chemistry Answers eBook Subscription Services
 - Review Acids And Bases Modern Chemistry Answers Budget-Friendly Options
6. Navigating Review Acids And Bases Modern Chemistry Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Review Acids And Bases Modern Chemistry Answers Compatibility with Devices
 - Review Acids And Bases Modern Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Acids And Bases Modern Chemistry Answers
 - Highlighting and Note-Taking Review Acids And Bases Modern Chemistry Answers
 - Interactive Elements Review Acids And Bases Modern Chemistry Answers
 8. Staying Engaged with Review Acids And Bases Modern Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Acids And Bases Modern Chemistry Answers
 9. Balancing eBooks and Physical Books Review Acids And Bases Modern Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Acids And Bases Modern Chemistry Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Review Acids And Bases Modern Chemistry Answers
 - Setting Reading Goals Review Acids And Bases Modern Chemistry Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Review Acids And Bases Modern Chemistry Answers
 - Fact-Checking eBook Content of Review Acids And Bases Modern Chemistry Answers
 - Distinguishing Credible Sources
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 - Exploring Educational eBooks
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