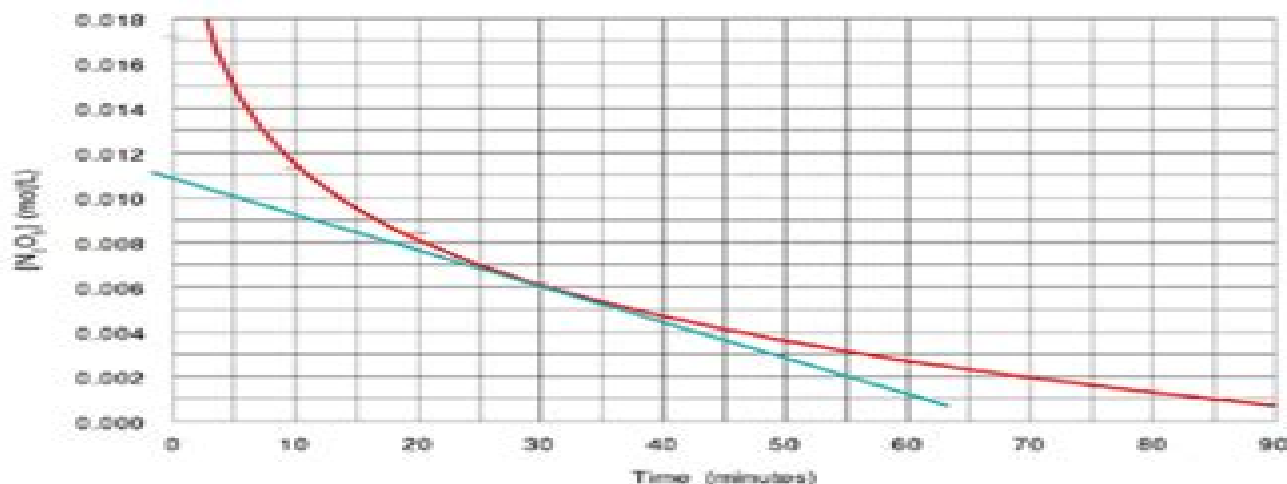


Kinetics Review Answers

1. Pressure, temperature, pH, conductivity, color, mass
2. To find the rate we need to calculate the slope of the tangent line at $t=30$ minutes.



$$\text{rate} = -\frac{\Delta[\text{N}_2\text{O}_5]}{\Delta t} = -\frac{(0.002 - 0.010)}{(55 - 5)} = 0.00016 \text{ mol/L min}$$

3. use the rate equation

a. $\frac{1}{2} \frac{\Delta[\text{N}_2]}{\Delta t} = \frac{1}{6} \frac{\Delta[\text{H}_2\text{O}]}{\Delta t}$

$$\frac{1}{2} (0.27) = \frac{1}{6} \frac{\Delta[\text{H}_2\text{O}]}{\Delta t}$$

$$\frac{\Delta[\text{H}_2\text{O}]}{\Delta t} = 0.81 \text{ mol/Ls}$$

b. $\frac{1}{2} \frac{\Delta[\text{N}_2]}{\Delta t} = \frac{1}{4} \frac{\Delta[\text{NH}_3]}{\Delta t}$

$$\frac{1}{2} (0.27) = \frac{1}{4} \frac{\Delta[\text{NH}_3]}{\Delta t}$$

$$\frac{\Delta[\text{NH}_3]}{\Delta t} = 0.54 \text{ mol/Ls}$$

c. $\frac{1}{2} \frac{\Delta[\text{N}_2]}{\Delta t} = \frac{1}{3} \frac{\Delta[\text{O}_2]}{\Delta t}$

$$\frac{1}{2} (0.27) = \frac{1}{3} \frac{\Delta[\text{O}_2]}{\Delta t}$$

$$\frac{\Delta[\text{O}_2]}{\Delta t} = 0.405 \text{ mol/Ls}$$

4. Surface area – an increased surface area provides more particles for collisions; more collisions means an increase in reaction rate

Concentration – increased concentration means more particles are present; more particles means more collisions which means an increase in reaction rate

Temperature – an increase in temperature increases the energy of the particles; increased energy means that more of the collisions will have sufficient energy to form the activated complex resulting in an increased reaction rate.

Review And Reinforcement Chemical Kinetics Answers

Ying Liu



Review And Reinforcement Chemical Kinetics Answers:

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