

CHAPTER 9 REVIEW

Stoichiometry

MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

1. Given the following equation:
- $\text{C}_3\text{H}_8(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$

_____ ⁴ a. What is the value of the coefficient x in this equation?_____ ^{40.0 g/mol} b. What is the molar mass of C_3H_8 ?_____ ^{4 mol O₂}
_____ ^{2 mol H₂O} c. What is the mole ratio of O_2 to H_2O in the above equation?_____ ^{0.20 mol} d. How many moles are in an 8.0 g sample of C_3H_8 ?_____ ³² e. If x mol of C_3H_8 react, how many moles of CO_2 are produced, in terms of x ? $\text{mol CO}_2 = (x \text{ mol}) \left(\frac{3 \text{ mol CO}_2}{1 \text{ C}_3\text{H}_8} \right) = 3x$

2. a. What is meant by
- ideal conditions*
- relative to stoichiometric calculations?

_____ ^{STP $\Rightarrow$ 101.3 kPa}
_____ ^{0°C}

- b. What function do ideal stoichiometric calculations serve?

_____ ^{allow us to use known/constant values}

- c. Are actual yields typically larger or smaller than theoretical yields?

_____ ^{smaller}**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

3. Assume the reaction represented by the following equation goes all the way to completion:



- _____
- ^{4 mol H₂}
- a. If 6 mol of
- H_2
- are consumed, how many moles of
- NH_3
- are produced?

$$\text{mol NH}_3 = (6 \text{ mol H}_2) \left(\frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) = 4 \text{ mol}$$

- _____
- ^{8.5 g NH₃}
- b. How many grams are in a sample of
- NH_3
- that contains
- 3.0×10^{23}
- molecules?

$$g \text{ NH}_3 = (3.0 \times 10^{23} \text{ molecules}) \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) \left(\frac{17.0 \text{ g}}{\text{mol}} \right)$$

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