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Unit 4: Gases Unit Review Sheet

Chemistry R

1. What are the 4 major ideas of Kinetic Molecular Theory?

- A) Random, straight line motion, constant
- B) elastic collisions
- C) volume of particles is insignificant
- D) no attractive forces, no IMF

2. Ideal vs. Real Gases

PLIGHT He, H₂

- no attractive forces
- light & small gases

3. Gas Laws

A) Boyle's: $P \downarrow, V \uparrow$ ($P \uparrow, V \downarrow$)

Formula: $P_1 V_1 = P_2 V_2$

- Marshmallow/Peep Experiment

($\downarrow P$ around the peep, so P inside was now greater than P outside of peep, so V of peep $\uparrow$, peep will expand)

B) Charles: $V \uparrow, T \uparrow$ ($T \downarrow, V \downarrow$)

Formula: _____

- > Balloon in the freezer example

$\downarrow$ $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

($\downarrow T$, the molecules slow down which means fewer collisions with the sides of the balloon, so not much pressure inside so $V \downarrow$)

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Rachel Sandford



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