

Chemistry Semester Exam Review

Complete all problems on notebook paper. Transfer your final answers to the answer sheet.

Measurements and Calculations

Significant Figures – adding, subtracting, multiplying, dividing

Scientific Notation - adding, subtracting, multiplying, dividing

Metric Conversions

Dimensional Analysis

Temperature Conversions

Density – $D=m/V$, $V=LWH$, $V=\pi r^2h$, $V=V_f - V_i$

Determine the number of significant figures in the following numbers.

- a. 1 SF b. 2 SF c. 4 SF d. 6 SF e. 10 SF

1. 0.0000787690 g
2. 900000 km
3. 0.009001 mm
4. 9.100 cg

Record your answer with the correct number of significant figures and units.

5. $45.6 \text{ g} + 3.59 \text{ g}$
a. 49.19 g b. 49.0 g c. 49.2 g d. 49.20 g
6. $0.00134 \text{ mL} - 0.00023 \text{ mL}$
a. 0.00111 mL b. 11 mL c. 0.0011 mL d. 0.00157 mL
7. $(56.2 \text{ cm})(1.234 \text{ cm})$
a. 69.3 cm^2 b. 69.4 cm^2 c. 69.3508 cm^2 d. 69.35 cm^2
8. $(3.20 \text{ km}^2)/(4.0 \text{ km})$
a. 13 km b. 8 km c. 0.80 km d. 2.56 km

Convert the following numbers from scientific notation to ordinary notation.

9. $3.02 \times 10^{-3} \text{ g}$
a. $.302 \times 10^{-3} \text{ g}$ b. 3.02 g c. 0.00302 g d. 3020 g
10. $5.791 \times 10^5 \text{ m}$
a. 579100 m b. 5.791 m c. 0.00005791 m d. 57910 m

Convert the following numbers from ordinary notation to scientific notation.

11. 4560 cm
a. $456 \times 10^3 \text{ cm}$ b. $4.56 \times 10^3 \text{ cm}$ c. $4.56 \times 10^{-3} \text{ cm}$ d. 4.56 cm
12. 0.0076 g
a. $7.6 \times 10^{-3} \text{ g}$ b. $7.6 \times 10^3 \text{ g}$ c. $76 \times 10^{-3} \text{ g}$ d. $0.0076 \times 10^{-3} \text{ g}$

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Xiaolong Qi



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